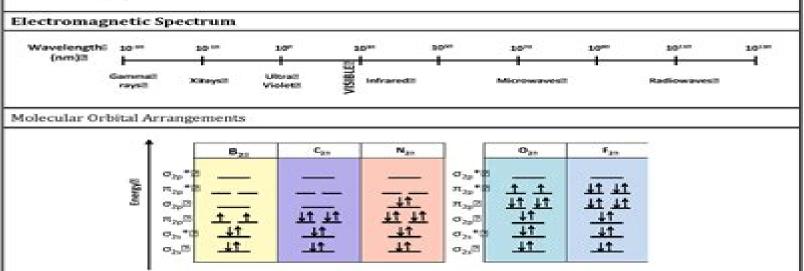
General Chemistry 1: Exam Appendix

Constants	Equations		
R = 8.314 J/mol-K R = 0.082 L-atm/mol-K	$E = (-2.18 \times 10^{-10} \text{J}) \text{ Z}^2/\text{n}^2$	$E = hc/\lambda$	$E = mc^2$
$h = 6.626 \times 10^{-34} \text{ J-s}$	$KE = (1/2)mv^2$	KE = (3/2)nRT	
$c = 3.00 \times 10^{8} \text{ m/s}$ $N_A = 6.022 \times 10^{23} \text{ particles/mole}$	PV = nRT	$T(K) = T(^{\circ}C) + 273$	

Solubility Rules

- All ammonium and group 1A salts are soluble
- All nitrates are soluble.
- All acetates, chlorates, and perchlorates are soluble
- Most chlorides, bromides, and iodides are soluble except silver, mercury, and lead
- Most sulfates are soluble except calcium, strontium, barium, and lead
- Phosphates are insoluble (except ammonium and Group 1A)
- Carbonates are insoluble (except ammonium and Group 1A)
- Chromates (CrO₄²·) are insoluble (except ammonium, Group 1A, and magnesium)
- Hydroxides are insoluble (except ammonium, Group 1A, barium, strontium, calcium)
- Sulfides are insoluble (except ammonium, Group 1A, magnesium, calcium, strontium, barium)



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