

# Algebra II Exam C Review One

key

Solve each equation. Be sure to check your solutions!

1.  $\frac{2}{7} + \frac{5}{x} = \frac{45}{7x}$   $x = 5$

$x \neq 0$

2.  $\frac{2}{x-1} = \frac{5}{x}$

$x \neq 0, 1$

$x = \frac{5}{3}$

3.  $\frac{14}{x} = 9 - x$   $x = 2, 7$

$x \neq 0$

4.  $\frac{x+5}{x-1} + \frac{x}{2} = \frac{12}{2x-2}$

$x \neq 1$

$x = -2$

5.  $\frac{2}{x-3} + \frac{5}{x+3} = \frac{7}{x^2-9}$

$x \neq 3, -3$

$x = \frac{14}{7}$

Simplify each expression. Assume all variables are positive.

6.  $\sqrt{x^2 - 12x + 36}$   $x - 6$

7.  $\sqrt{32x^3}$   $4x\sqrt{2}$

Write in radical form and simplify.

8.  $16^{\frac{5}{4}}$  Radical form:  $\sqrt[4]{16^5}$  Simplified:  $32$

Write using rational exponents. Then simplify.

9.  $(\sqrt[3]{-1000})^2$  Rational exponents:  $(-1000)^{\frac{2}{3}}$  Simplified:  $100$

Simplify.

10.  $\sqrt{m} + \sqrt{m^3}$   $(1+m)\sqrt{m}$

11.  $2^{\frac{1}{2}} - 3\sqrt{2}$   $-2\sqrt[3]{2}$

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**Sabine Zange**



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web the volume of 1 00 mol 1 00 mol of any gas at stp standard temperature 273 15 k and pressure 1 atm is measured to be 22 414l 22 414 l we can substitute 101 325kpa 101 325 kpa for pressure 22 414 l 22 414 l for volume and 273 15 k 273 15 k for temperature into the ideal gas equation and solve for r r

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web chemists sometimes make comparisons against a standard temperature and pressure stp for reporting properties of gases 273 15 k and 1 atm 101 325 kpa 1 at stp one mole of an ideal gas has a volume of about 22 4 l this is referred to as the standard molar volume figure 9 18

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web volume pressure kpa linear pressure kpa volume ml pressure kpa in part b we observed the relationship between temperature and pressure the pressure of the 125 ml flask was at 101 36 kpa when the temperature was at 23 2 c

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web pressure increases even though there were errors in the experiment the goal of the experiment was met the goal of declaring what type of relationship it is was clear using the data pressure kpa temperature c temperature k constant k p t 95 82 kpa 16 0 c 289 k 0 33 97 90 kpa 19 0 c 292 k 0 34 98 97 kpa 99 22 kpa 101 4 kpa 19 1 c

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web the properties of water at constant volume can be represented as a function of pressure and temperature as shown in the diagram below figure 1 the pressure temperature relationship at constant volume the saturation point of water is the condition at which a phase change occurs from liquid to vapour or vapour to liquid

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web figure pageindex 5 the empirically determined relationships among pressure volume temperature and amount of a gas the thermometer and pressure gauge indicate the temperature and the pressure qualitatively the level in the flask indicates the volume and the number of particles in each flask indicates relative amounts

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