

The Synthesis of Biodiesel from Vegetable Oil

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Introduction

The use of reclaimed vegetable oil from restaurants, for use as a fuel for road vehicles, has received a lot of attention in recent years. Used vegetable oils contain solids and free fatty acids due to oil breakdown during the frying process. Normally, the used oil must be filtered and the quantity of the free fatty acids must be determined so that the chemicals necessary for the breaking down of the cooking oil can be calculated.

In this experiment the biodiesel will be made from fresh vegetable oil to avoid the solids and a titration to determine the mass of sodium hydroxide needed to react with the whole oil molecules.

The diesel engine was developed by Rudolf Diesel, a French thermal engineer, in 1893 with the first working prototype in 1897. A diesel engine works by creating heat by compressing air followed by the injection of fuel which burns, increasing pressure to drive the pistons. Early diesel engines were intended to run on peanut oil and were later adapted to run on the lower viscosity and cheaper petroleum diesel. Today, diesel engines are used in cars, buses and trucks. Biodiesel can be used directly in diesel engines or blended with petroleum based diesel fuel.

Vegetable oils are esters of glycerin, commonly called triglycerides, with different fatty acids with the structure:



Where R^I , R^{II} , and R^{III} can be the same or different fatty acids. (See Table 1)

The preparation of the biodiesel is a transesterification reaction where the triglycerides are converted into simpler methyl esters of the fatty acids (the biodiesel).



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Mike Jess

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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