

Задача 11.
Слабые и сред. кислоты

① Дано:

$$C_{\text{HNO}_2} = 0,01 \text{ м}$$

$$K_a = 7,9 \cdot 10^{-4}$$

$$pH = ?$$

$$pH = -\lg [H^+] = -\lg 2,8 \cdot 10^{-5} = 4,55$$

Решение:

Для слабой кислоты:

$$[H^+] = \sqrt{K_a \cdot C_{\text{HNO}_2}}$$

$$= \sqrt{7,9 \cdot 10^{-4} \cdot 10^{-2}} = 2,8 \cdot 10^{-5} \text{ моль/л}$$

$$= 4,55$$

② Дано:

$$K_{\text{PrOH}_2} = 6,3 \cdot 10^{-5}$$

$$pH = ?$$

Решение:

$$C_{\text{PrOH}_2} = C_{\text{H}^+} = C_{\text{OH}^-}$$

$$K_{\text{Pr}} = [C_{\text{H}^+}][C_{\text{OH}^-}] = 6,3 \cdot 10^{-5}$$

S - концентрация ионов

$$S = \sqrt[3]{\frac{K_{\text{Pr}}}{4}} = \sqrt[3]{\frac{6,3 \cdot 10^{-5}}{4}} = 1,16 \cdot 10^{-2} \text{ моль/л}$$

$$[C_{\text{OH}^-}] = \frac{S}{2} = \frac{1,16 \cdot 10^{-2}}{2} = 5,8 \cdot 10^{-3} \text{ моль/л}$$

$$[C_{\text{H}^+}] = \frac{10^{-14}}{[C_{\text{OH}^-}]} = \frac{10^{-14}}{5,8 \cdot 10^{-3}} = 1,72 \cdot 10^{-12}$$

$$pH = -\lg [C_{\text{H}^+}] = -\lg 1,72 \cdot 10^{-12} = 11,76$$

③ Дано:

$$V_{\text{HNO}_2} = V_{\text{NaOH}}$$

$$C_{\text{HNO}_2} = 0,1 \text{ м}$$

$$V_{\text{PrOH}_2} = 1 \text{ см}^3$$

$$C_{\text{PrOH}_2} = 0,01 \text{ м}$$

$$K_{\text{PrOH}_2} = 1,6 \cdot 10^{-5}$$

Решение:

полн. самодиссоции:

$$C_{\text{HNO}_2} = 0,1 \frac{\text{г}}{\text{л}} = 0,05 \text{ моль/л}$$

$$C_{\text{PrOH}_2} = 0,01 \frac{\text{г}}{\text{л}} = 0,005 \text{ моль/л}$$

Для диссоциирующей кислоты:

$$C_{\text{H}^+} = C_{\text{HNO}_2} = 0,05 \text{ моль/л}$$

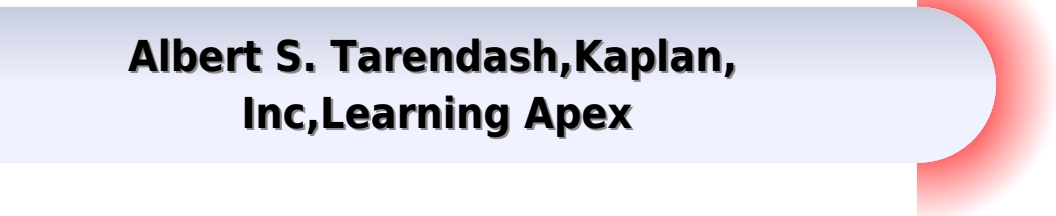
$$C_{\text{PrOH}_2} = C_{\text{PrOH}_2} = 0,005 \text{ моль/л}$$

$$[C_{\text{Pr}^+}][C_{\text{OH}^-}] = 1,6 \cdot 10^{-5}$$

$$PrOH_2 \rightleftharpoons Pr^+ + OH^-$$

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**Albert S. Tarendash, Kaplan,
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