

Задача 11.
Слабые и сред. кислоты

① Дано:

$$C_{\text{HAc}} = 0,01 \text{ м}$$

$$K_a = 7,9 \cdot 10^{-6}$$

$$pH = ?$$

$$pH = -\lg [H^+] = -\lg 2,8 \cdot 10^{-6} = 6 - \lg 2,8 = 6 - 0,45 = 5,55$$

Решение:

Для слабой кислоты:

$$[H^+] = \sqrt{K_a \cdot C_{\text{HAc}}}$$

$$= \sqrt{7,9 \cdot 10^{-6} \cdot 10^{-2}} = 2,8 \cdot 10^{-6} \text{ моль/л}$$

② Дано:

$$K_{\text{PrOH}_2} = 6,3 \cdot 10^{-5}$$

$$pH = ?$$

Решение:

$$C_{\text{PrOH}_2} = C_{\text{H}^+} = C_{\text{OH}^-} = 2,5 \cdot 10^{-3}$$

$$K_{\text{Pr}} = [C_{\text{H}^+}][\text{OH}^-] = 2,5 \cdot 10^{-3} \cdot 2,5 \cdot 10^{-3} = 6,25 \cdot 10^{-6}$$

S - концентрация ионов

$$S = \sqrt[3]{\frac{K_{\text{Pr}}}{4}} = \sqrt[3]{\frac{6,25 \cdot 10^{-6}}{4}} = 1,16 \cdot 10^{-2} \text{ моль/л}$$

$$[OH^-] = \frac{S}{2} = \frac{1,16 \cdot 10^{-2}}{2} = 5,8 \cdot 10^{-3} \text{ моль/л}$$

$$[H^+] = \frac{10^{-14}}{[OH^-]} = \frac{10^{-14}}{5,8 \cdot 10^{-3}} = 1,72 \cdot 10^{-12}$$

$$pH = -\lg [H^+] = -\lg 1,72 \cdot 10^{-12} = 12 - \lg 1,72 = 11,76$$

③ Дано:

$$V_{\text{HAc}} + V_{\text{NaOH}}$$

$$C_{\text{HAc}} = 0,1 \text{ м}$$

$$V_{\text{PrOH}_2} = 10 \text{ мл}$$

$$C_{\text{PrOH}_2} = 0,01 \text{ м}$$

$$K_{\text{PrOH}_2} = 6,3 \cdot 10^{-5}$$

Решение:

полн. самодиссоции:

$$C_{\text{HAc}} = 0,1 \cdot \frac{10}{20} = 0,05 \text{ моль/л}$$

$$C_{\text{PrOH}_2} = 0,01 \cdot \frac{10}{20} = 0,005 \text{ моль/л}$$

или диссоциирует комплекс

$$C_{\text{HAc}} = C_{\text{HAc}} = 0,05 \text{ моль/л}$$

$$C_{\text{PrOH}_2} = C_{\text{PrOH}_2} = 0,005 \text{ моль/л}$$

$$PrOH_2 \rightleftharpoons PrO^- + H^+$$

$$K_{\text{PrOH}_2} = 6,3 \cdot 10^{-5}$$

$$[PrO^-][H^+] = 6,3 \cdot 10^{-5} \cdot 0,005 = 3,15 \cdot 10^{-7}$$

Apex Chemistry Semester 2 Exam Answers

Rob Franek



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