

AP CHEMISTRY

Volume 1: Your Second Year Study

Properties of Solids

Crystalline Solids

- Have a regular, repeating arrangement of particles (atoms, ions, or molecules).
- Have sharp melting points.
- Can be cleaved along flat planes.

Amorphous Solids

- Do not have a regular arrangement of particles.
- Do not have sharp melting points.
- Do not cleave along flat planes.

Intermolecular Forces

- London Dispersion Forces:** Present in all molecules, caused by temporary dipoles.
- Dipole-Dipole Forces:** Present in polar molecules.
- Hydrogen Bonding:** A special type of dipole-dipole force between molecules containing N, O, or F.

Periodic Table & Trends

Electronegativity (increases from bottom-left to top-right)

Ionization Energy (increases from bottom-left to top-right)

Atomic Radius (increases from top-left to bottom-right)

Properties of the Periodic Table

- Metals:** Located on the left side of the table. They are shiny, malleable, and good conductors of heat and electricity.
- Nonmetals:** Located on the right side of the table. They are not shiny, brittle, and poor conductors.
- Metalloids:** Located between metals and nonmetals. They have properties of both.
- Alkali Metals:** Group 1. Very reactive, soft, and have low melting points.
- Alkaline Earth Metals:** Group 2. More reactive than alkali metals.
- Halogens:** Group 17. Very reactive, nonmetals.
- Noble Gases:** Group 18. Inert, no chemical reactions.

Electronegativity Table

Electronegativity increases from bottom-left to top-right.

1.0	2.1	2.5	3.0	3.5	4.0	4.5	5.0	5.5	6.0	6.5	7.0	7.5	8.0	8.5	9.0	9.5	10.0
1.1	2.2	2.6	3.1	3.6	4.1	4.6	5.1	5.6	6.1	6.6	7.1	7.6	8.1	8.6	9.1	9.6	10.1
1.2	2.3	2.7	3.2	3.7	4.2	4.7	5.2	5.7	6.2	6.7	7.2	7.7	8.2	8.7	9.2	9.7	10.2
1.3	2.4	2.8	3.3	3.8	4.3	4.8	5.3	5.8	6.3	6.8	7.3	7.8	8.3	8.8	9.3	9.8	10.3
1.4	2.5	2.9	3.4	3.9	4.4	4.9	5.4	5.9	6.4	6.9	7.4	7.9	8.4	8.9	9.4	9.9	10.4
1.5	2.6	3.0	3.5	4.0	4.5	5.0	5.5	6.0	6.5	7.0	7.5	8.0	8.5	9.0	9.5	10.0	10.5
1.6	2.7	3.1	3.6	4.1	4.6	5.1	5.6	6.1	6.6	7.1	7.6	8.1	8.6	9.1	9.6	10.1	10.6
1.7	2.8	3.2	3.7	4.2	4.7	5.2	5.7	6.2	6.7	7.2	7.7	8.2	8.7	9.2	9.7	10.2	10.7
1.8	2.9	3.3	3.8	4.3	4.8	5.3	5.8	6.3	6.8	7.3	7.8	8.3	8.8	9.3	9.8	10.3	10.8
1.9	3.0	3.4	3.9	4.4	4.9	5.4	5.9	6.4	6.9	7.4	7.9	8.4	8.9	9.4	9.9	10.4	10.9
2.0	3.1	3.5	4.0	4.5	5.0	5.5	6.0	6.5	7.0	7.5	8.0	8.5	9.0	9.5	10.0	10.5	11.0

Types of Chemical Calculations

Stoichiometry: Calculations involving the relative amounts of reactants and products in a chemical reaction.

Limiting Reactant: The reactant that is completely consumed in a reaction, limiting the amount of product that can be formed.

Percent Yield: The ratio of the actual yield to the theoretical yield, expressed as a percentage.

Gas Laws: Calculations involving the pressure, volume, and temperature of a gas.

Solubility: Calculations involving the amount of a substance that can dissolve in a solvent.

Acids & Bases

Acids: Substances that donate protons (H⁺).

Bases: Substances that accept protons (H⁺).

Strong Acids: Completely dissociate in water.

Weak Acids: Partially dissociate in water.

Strong Bases: Completely dissociate in water.

Weak Bases: Partially dissociate in water.

pH Scale: A measure of the acidity or basicity of a solution, ranging from 0 to 14.

Acid-Base Reactions

Acid	Base	Conjugate Acid	Conjugate Base
HCl	NaOH	NaCl	H ₂ O
HNO ₃	KOH	KNO ₃	H ₂ O
H ₂ SO ₄	Ca(OH) ₂	CaSO ₄	2H ₂ O
H ₃ PO ₄	NaOH	Na ₃ PO ₄	3H ₂ O
CH ₃ COOH	NaOH	CH ₃ COONa	H ₂ O
H ₂ CO ₃	NaOH	Na ₂ CO ₃	2H ₂ O
H ₂ SO ₃	NaOH	Na ₂ SO ₃	2H ₂ O
H ₂ SiO ₃	NaOH	Na ₂ SiO ₃	2H ₂ O
H ₂ BO ₃	NaOH	Na ₂ BO ₃	2H ₂ O
H ₂ GeO ₃	NaOH	Na ₂ GeO ₃	2H ₂ O
H ₂ SnO ₃	NaOH	Na ₂ SnO ₃	2H ₂ O
H ₂ PbO ₃	NaOH	Na ₂ PbO ₃	2H ₂ O
H ₂ UO ₃	NaOH	Na ₂ UO ₃	2H ₂ O
H ₂ ThO ₃	NaOH	Na ₂ ThO ₃	2H ₂ O
H ₂ PaO ₃	NaOH	Na ₂ PaO ₃	2H ₂ O
H ₂ AcO ₃	NaOH	Na ₂ AcO ₃	2H ₂ O
H ₂ AsO ₃	NaOH	Na ₂ AsO ₃	2H ₂ O
H ₂ SbO ₃	NaOH	Na ₂ SbO ₃	2H ₂ O
H ₂ BiO ₃	NaOH	Na ₂ BiO ₃	2H ₂ O
H ₂ AtO ₃	NaOH	Na ₂ AtO ₃	2H ₂ O
H ₂ FrO ₃	NaOH	Na ₂ FrO ₃	2H ₂ O
H ₂ BrO ₃	NaOH	Na ₂ BrO ₃	2H ₂ O
H ₂ IO ₃	NaOH	Na ₂ IO ₃	2H ₂ O
H ₂ TeO ₃	NaOH	Na ₂ TeO ₃	2H ₂ O
H ₂ SeO ₃	NaOH	Na ₂ SeO ₃	2H ₂ O
H ₂ SO ₃	NaOH	Na ₂ SO ₃	2H ₂ O
H ₂ SiO ₃	NaOH	Na ₂ SiO ₃	2H ₂ O
H ₂ BO ₃	NaOH	Na ₂ BO ₃	2H ₂ O
H ₂ GeO ₃	NaOH	Na ₂ GeO ₃	2H ₂ O
H ₂ SnO ₃	NaOH	Na ₂ SnO ₃	2H ₂ O
H ₂ PbO ₃	NaOH	Na ₂ PbO ₃	2H ₂ O
H ₂ UO ₃	NaOH	Na ₂ UO ₃	2H ₂ O
H ₂ ThO ₃	NaOH	Na ₂ ThO ₃	2H ₂ O
H ₂ PaO ₃	NaOH	Na ₂ PaO ₃	2H ₂ O
H ₂ AcO ₃	NaOH	Na ₂ AcO ₃	2H ₂ O
H ₂ AsO ₃	NaOH	Na ₂ AsO ₃	2H ₂ O
H ₂ SbO ₃	NaOH	Na ₂ SbO ₃	2H ₂ O
H ₂ BiO ₃	NaOH	Na ₂ BiO ₃	2H ₂ O
H ₂ AtO ₃	NaOH	Na ₂ AtO ₃	2H ₂ O
H ₂ FrO ₃	NaOH	Na ₂ FrO ₃	2H ₂ O
H ₂ BrO ₃	NaOH	Na ₂ BrO ₃	2H ₂ O
H ₂ IO ₃	NaOH	Na ₂ IO ₃	2H ₂ O
H ₂ TeO ₃	NaOH	Na ₂ TeO ₃	2H ₂ O
H ₂ SeO ₃	NaOH	Na ₂ SeO ₃	2H ₂ O
H ₂ SO ₃	NaOH	Na ₂ SO ₃	2H ₂ O
H ₂ SiO ₃	NaOH	Na ₂ SiO ₃	2H ₂ O
H ₂ BO ₃	NaOH	Na ₂ BO ₃	2H ₂ O
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H ₂ PbO ₃	NaOH	Na ₂ PbO ₃	2H ₂ O
H ₂ UO ₃	NaOH	Na ₂ UO ₃	2H ₂ O
H ₂ ThO ₃	NaOH	Na ₂ ThO ₃	2H ₂ O
H ₂ PaO ₃	NaOH	Na ₂ PaO ₃	2H ₂ O
H ₂ AcO ₃	NaOH	Na ₂ AcO ₃	2H ₂ O
H ₂ AsO ₃	NaOH	Na ₂ AsO ₃	2H ₂ O
H ₂ SbO ₃	NaOH	Na ₂ SbO ₃	2H ₂ O
H ₂ BiO ₃	NaOH	Na ₂ BiO ₃	2H ₂ O
H ₂ AtO ₃	NaOH	Na ₂ AtO ₃	2H ₂ O
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H ₂ BO ₃	NaOH	Na ₂ BO ₃	2H ₂ O
H ₂ GeO ₃	NaOH	Na ₂ GeO ₃	2H ₂ O
H ₂ SnO ₃	NaOH	Na ₂ SnO ₃	2H ₂ O
H ₂ PbO ₃	NaOH	Na ₂ PbO ₃	2H ₂ O
H ₂ UO ₃	NaOH	Na ₂ UO ₃	2H ₂ O
H ₂ ThO ₃	NaOH	Na ₂ ThO ₃	2H ₂ O
H ₂ PaO ₃	NaOH	Na ₂ PaO ₃	2H ₂ O
H ₂ AcO ₃	NaOH	Na ₂ AcO ₃	2H ₂ O
H ₂ AsO ₃	NaOH	Na ₂ AsO ₃	2H ₂ O
H ₂ SbO ₃	NaOH	Na ₂ SbO ₃	2H ₂ O
H ₂ BiO ₃	NaOH	Na ₂ BiO ₃	2H ₂ O
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H ₂ SiO ₃	NaOH	Na ₂ SiO ₃	2H ₂ O
H ₂ BO ₃	NaOH	Na ₂ BO ₃	2H ₂ O
H ₂ GeO ₃	NaOH	Na ₂ GeO ₃	2H ₂ O
H ₂ SnO ₃	NaOH	Na ₂ SnO ₃	2H ₂ O
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