

CHAPTER 7

STUDY GUIDE

Ionic Compounds and Metals

Section 7.1 Ion Formation

In your textbook, read about chemical bonds and formation of ions.

Use each of the terms below just once to complete the passage.

chemical bond	electrons	energy level	ions	noble gases
nucleus	octet	pseudo-noble gas formations		valence

The force that holds two atoms together is called a(n) (1) _____.

Such an attachment may form by the attraction of the positively charged

(2) _____ of one atom for the negatively charged

(3) _____ of another atom, or by the attraction of charged atoms, which are called

(4) _____. The attractions may also involve

(5) _____ electrons, which are the electrons in the outermost

(6) _____. The (7) _____ are a family of elements that have very

little tendency to react. Most of these elements have a set of eight outermost electrons, which is called a

stable (8) _____. The relatively stable electron structures developed by loss of

electrons in certain elements of groups 3, 4, 13, and 14 are called (9) _____.

For each statement below, write *true* or *false*.

- _____ 10. A positively charged ion is called an anion.
- _____ 11. Elements in group 1 lose their one valence electron, forming an ion with a 1+ charge.
- _____ 12. Elements tend to react so that they acquire the electron structure of a halogen.
- _____ 13. A sodium atom tends to lose one electron when it reacts.
- _____ 14. The electron structure of a zinc ion (Zn^{2+}) is an example of a pseudo-noble gas formation.
- _____ 15. A Cl^- ion is an example of a cation.
- _____ 16. The ending *-ide* is used to designate an anion.
- _____ 17. Nonmetals form a stable outer electron configuration by losing electrons and becoming anions.

Answer Study Guide Ionic Compounds And Metals

Pamela Mills,Amina El-Ashmawy



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Table of Contents Answer Study Guide Ionic Compounds And Metals

1. Understanding the eBook Answer Study Guide Ionic Compounds And Metals
 - The Rise of Digital Reading Answer Study Guide Ionic Compounds And Metals
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Study Guide Ionic Compounds And Metals
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Study Guide Ionic Compounds And Metals
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Study Guide Ionic Compounds And Metals
 - Personalized Recommendations
 - Answer Study Guide Ionic Compounds And Metals User Reviews and Ratings
 - Answer Study Guide Ionic Compounds And Metals and Bestseller Lists
5. Accessing Answer Study Guide Ionic Compounds And Metals Free and Paid eBooks
 - Answer Study Guide Ionic Compounds And Metals Public Domain eBooks
 - Answer Study Guide Ionic Compounds And Metals eBook Subscription Services
 - Answer Study Guide Ionic Compounds And Metals Budget-Friendly Options
6. Navigating Answer Study Guide Ionic Compounds And Metals eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Study Guide Ionic Compounds And Metals Compatibility with Devices
 - Answer Study Guide Ionic Compounds And Metals Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Study Guide Ionic Compounds And Metals
 - Highlighting and Note-Taking Answer Study Guide Ionic Compounds And Metals
 - Interactive Elements Answer Study Guide Ionic Compounds And Metals

8. Staying Engaged with Answer Study Guide Ionic Compounds And Metals
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Study Guide Ionic Compounds And Metals
9. Balancing eBooks and Physical Books Answer Study Guide Ionic Compounds And Metals
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Study Guide Ionic Compounds And Metals
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Study Guide Ionic Compounds And Metals
 - Setting Reading Goals Answer Study Guide Ionic Compounds And Metals
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Study Guide Ionic Compounds And Metals
 - Fact-Checking eBook Content of Answer Study Guide Ionic Compounds And Metals
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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