## Basic Transport Phenomena in Biomedical Engineering 4th Edition Fournier Solutions Manual

#### Chapter I Solutions

Problem 1.1. Use the conversion factors

$$\frac{8.314J}{\text{mol K}} \times \frac{0.23901\text{cal}}{J} = 1.987 \frac{\text{cal}}{\text{mol K}}$$

**Problem 1.2.** Use Equation 1.2 to convert the temperature to degrees Celsius

$$t \cdot C = \frac{5}{9} (t \cdot F - 32) = \frac{5}{9} (98.6 - 32) = 37 \cdot C$$

Now, use Equation 1.3 to convert the temperature from degrees Celsius to Kelvin

$$TK = t \cdot C + 273.15 = 37 + 273.15 = 310.15 K$$

**Problem 1.3.** Use the conversion factors

$$2 \times 10^4 \,\mu dynes \times \frac{dynes}{10^6 \,\mu dynes} \times \frac{N}{105 dynes} \times \frac{kN}{10^3 \,N} = 2 \times 10^{-10} \,kN$$

Problem 1.4. Use the conversion factors

$$\frac{27\,\mu m}{hr} \times \frac{10^{-6} \text{ m}}{\mu m} \times \frac{0.0006214 \text{miles}}{m} \times \frac{hr}{60 \text{min}} \times \frac{min}{60 \text{sec}} = 4.66 \times 10^{-12} \frac{\text{miles}}{\text{sec}}$$

Problem 1.5. Use the conversion factors

$$48MW \times \frac{106W}{MW} \times \frac{kg \text{ m}^2}{\text{sec}^3 \text{ W}} \times \frac{10^3 \text{ g}}{\text{kg}} \times \frac{100 \text{cg}}{\text{g}} \times \frac{3600^3 \text{ sec}^3}{\text{hr}^5} \times \frac{100^2 \text{ cm}^2}{\text{m}^2} = 2.24 \times 10^{27} \frac{\text{cgcm}^2}{\text{hr}^3}$$

$$48MW \times \frac{10^6 \text{ W}}{MW} \times \frac{J}{\text{sec W}} \times \frac{kJ}{10^3 \text{ J}} \times \frac{60 \text{sec}}{-\text{min}} = 2.88 \times 10^6 \frac{kJ}{\text{min}}$$

Problem 1.6. Use the conversion factors to get both values into units of mm per second and then compare the two values for the higher value.

# Basic Transport Phenomena In Biomedical Engineeringthird Edition

**Ronald L. Fournier** 

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