

Some Basic Concept of Chemistry

1. Number of molecules in $W(g)$ of substance = $\frac{W(g) \times N_A}{GMM}$
2. Molality (m) = $\frac{\text{No. of moles of solute}}{\text{Mass of solvent in kg}}$
3. Number of molecules in V litre of gas at S.T.P. = $\frac{VN_A}{22.4}$
4. Number of gram atoms = $\frac{W(g)}{GAM}$ (GAM $\rightarrow$ gram atomic mass)
5. Number of gram molecules = $\frac{W(g)}{\text{Gram molecular mass}}$
6. Dilution formula: $M_1V_1 = M_2V_2$
For mixing two solutions of the same substance
 $M_1V_1 + M_2V_2 = M_3(V_1 + V_2)$
Molarity can be directly calculated from % by mass (w/w) if density is known
Molarity = $\frac{\% \times 10 \times d}{GMM}$
7. Mass of 1 atom of element = $\frac{GAM}{N_A}$
8. Mass of 1 molecule of substance = $\frac{MM}{N_A}$ (MM $\rightarrow$ Molar mass)
9. $T(K) = T(^{\circ}C) + 273.15$
10. Relative atomic mass = $\frac{\text{Mass of an atom of the element}}{\frac{1}{12} \times \text{Mass of an atom of carbon (C-12)}}$
11. Number of molecules in n moles of substance = $n \times N_A$
12. Mass % of an element in a compound = $\frac{\text{Mass of that element in 1 mole of the compound}}{\text{Molar mass of the compound}}$
13. Mass percent = $\frac{\text{Mass of solute}}{\text{Mass of solution}} \times 100$
14. $\frac{X_a}{1 - X_a} = \frac{\text{Molality} \times M_a}{1000}$ where M_a = mass of solvent
15. Molarity (M) = $\frac{\text{No. of moles of solute}}{\text{Volume of solution in litres}}$ mole / L.
16. Avogadro's No. $N_A = 6.022 \times 10^{23}$
17. $T(^{\circ}F) = \frac{9}{5} T(^{\circ}C) + 32$
18. Molecular mass = $2 \times \text{vapour density}$
19. Mole fraction of A = $\frac{\text{No. of moles of A}}{\text{No. of moles of solution}}$

Basic Physical Chemistry Second Edition

Raymond Chang, Charles M. Lovett



Basic Physical Chemistry Second Edition:

Basic Physical Chemistry for the Atmospheric Sciences Peter V. Hobbs, 2000-09-04 Revised and updated in 2000 Basic Physical Chemistry for the Atmospheric Sciences provides a clear concise grounding in the basic chemical principles required for studies of atmospheres oceans and earth and planetary systems Undergraduate and graduate students with little formal training in chemistry can work through the chapters and the numerous exercises within this book before accessing the standard texts in the atmospheric chemistry geochemistry and the environmental sciences The book covers the fundamental concepts of chemical equilibria chemical thermodynamics chemical kinetics solution chemistry acid and base chemistry oxidation reduction reactions and photochemistry In a companion volume entitled Introduction to Atmospheric Chemistry 2000 Cambridge University Press Peter Hobbs provides an introduction to atmospheric chemistry itself including its applications to air pollution acid rain the ozone hole and climate change Together these two books provide an ideal introduction to atmospheric chemistry for a variety of disciplines

Physical Chemistry for the Biosciences, second edition Raymond Chang, Charles M. Lovett, 2025-02-25 Physical Chemistry for the Biosciences has been optimized for a one semester course in physical chemistry for students of biosciences or a course in biophysical chemistry Most students enrolled in this course have taken general chemistry organic chemistry and a year of physics and calculus Fondly known as Baby Chang this best selling text is ack in an updated second edition for the one semester physical chemistry course Carefully crafted to match the needs and interests of students majoring in the life sciences Physical Chemistry for the Biosciences has been revised to provide students with a sophisticated appreciation for physical chemistry as the basis for a variety of interesting biological phenomena Major changes to the new edition include Discussion of intermolecular forces in chapter Detailed discussion of protein and nucleic acid structure providing students with the background needed to fully understand the biological applications of thermodynamics and kinetics described later in the book Expanded and updated descriptions of biological examples such as protein misfolding diseases photosynthesis and vision

Soil Physical Chemistry Donald L. Sparks, 2018-02-06 Soil Physical Chemistry Second Edition takes up where the last edition left off With comprehensive and contemporary discussions on equilibrium and kinetic aspects of major soil chemical process and reactions this excellent text reference presents new chapters on precipitation dissolution modeling of adsorption reactions at the mineral water interface and the chemistry of humic substances An emphasis is placed on understanding soil chemical reactions from a microscopic point of view and rigorous theoretical developments such as the use of modern in situ surface chemical probes such as x ray adsorption fine structure XAFS Fournier transform infrared FTIR spectroscopies and scanning probe microscopies SPM are discussed

Basic Physical Chemistry: The Route To Understanding (Revised Edition) E Brian Smith, 2013-06-03 This elegant book provides a student friendly introduction to the subject of physical chemistry It is by the author of the very successful Basic Chemical Thermodynamics and is written in the same well received popular style It is concise and more

compact than standard textbooks on the subject and emphasises the two important topics underpinning the subject quantum mechanics and the second law of thermodynamics Both topics are challenging to students because they focus on uncertainty and probability The book explains these fundamental concepts clearly and shows how they offer the key to understanding a wide range of chemical phenomena including atomic and molecular spectra the structure and properties of solids liquids and gases chemical equilibrium and the rates of chemical reactions This revised edition has enabled improvements and corrections to be made

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Physical Chemistry of Metallurgical Processes, Second Edition Mohammad Shamsuddin, 2021-06-22 This updated second edition retains its classroom tested treatment of physical chemistry of metallurgical topics such as roasting of sulfide minerals matte smelting converting structure properties and theories of slag reduction of oxides and reduction smelting interfacial phenomena steelmaking secondary steelmaking role of halides in extraction of metals refining hydrometallurgy and electrometallurgy and adds new data in worked out examples as well as up to date references to the literature The book further explains the physical chemistry of various metallurgical topics steps involved in extraction of metals such as roasting matte smelting converting reduction smelting steelmaking reactions deoxidation stainless steelmaking vacuum degassing refining leaching chemical precipitation ion exchange solvent extraction cementation gaseous reduction and electrowinning Each topic is illustrated with appropriate examples of applications of the technique in extraction of some common reactive rare or refractory metal together with worked out problems explaining the principle of the operation The problems require imagination and critical analyses and also encourage readers for creative application of thermodynamic data in metal extraction Updates and condenses text throughout the book by sequential arrangement of paragraphs in different chapters Maximizes readers understanding of the physicochemical principles involved in extraction production of common and rare reactive metals by pyro as well as hydrometallurgical routes Reinforces concepts presented with worked examples in each chapter explaining the process steps Explains the physical chemistry of various metallurgical steps such as roasting matte smelting converting and reduction smelting steelmaking aqueous processing etc in extraction of metals Collects and uniformly presents scattered information on physicochemical principles of metal production from various books and journals

Physical Properties of Materials, Second Edition Mary Anne White, 2011-06-28 Designed for advanced undergraduate students Physical Properties of Materials Second Edition

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Concise Physical Chemistry Donald W. Rogers, 2011-03-01 This book is a physical chemistry textbook that presents the essentials of physical chemistry as a logical sequence from its most modest beginning to contemporary research topics Many books currently on the market focus on the problem sets with a cursory treatment of the conceptual background and theoretical material whereas this book is concerned only with the conceptual development of the subject Comprised of 19 chapters the book will address ideal gas laws real gases the thermodynamics of simple systems thermochemistry entropy and the second law the Gibbs free energy equilibrium statistical approaches to thermodynamics the phase rule chemical kinetics liquids and solids solution chemistry conductivity electrochemical cells atomic theory wave mechanics of simple systems molecular orbital theory experimental determination of molecular structure and photochemistry and the theory of chemical kinetics

Physical Chemistry William M. Davis, 2011-12-06 Designed specifically for a two semester introductory course sequence in physical chemistry this text presents core principles and topics Straightforward and streamlined it presents the necessary amount of detail for comprehension Organized in such a way that the various topics covered are connected to each other it allows students to see physical chemistry as an interconnected discipline and not a series of unrelated concepts Each chapter in this new edition has been thoroughly updated and includes new information on computational applications more end of chapter problems and new chapters on nanotechnology and surface chemistry

Principles of Physical Chemistry Hans Kuhn, Horst-Dieter Försterling, David H. Waldeck, 2009-03-17 Principles of Physical Chemistry Second Edition uniquely uses simple physical models as well as rigorous treatments for understanding molecular and supramolecular systems and processes In this way the presentation assists students in developing an intuitive understanding of the subjects as well as skill in quantitative

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