

① $x = 0.01$ Cr, $Ni = 0.04$

Mg = $0.5 \times 0.95 = 0.475$ M (Mg nitrate)

$$0.475 = \frac{n}{0.05} \Rightarrow n = 0.02375 \text{ mol}$$

$$\therefore w = 6.0895 \text{ g}$$

Mg nitrate

Ni = $0.5 \times 0.04 = 0.02$ M

$$0.02 = \frac{n}{0.05} \Rightarrow n = 0.001 \text{ mol}$$

$$w = 0.2908 \text{ g}$$

Ni nitrate

Cr = $0.5 \times 0.01 = 0.005$

$$0.005 = \frac{n}{0.05} \Rightarrow n = 0.00025 \text{ mol}$$

$$w = 0.100035 \text{ g}$$

Cr nitrate

② $x = 0.03$ Cr, $Ni = 0.02$

Mg $w = 6.0895 \text{ g}$

Ni = $0.5 \times 0.02 = 0.01$ M

$$\therefore n = 0.0005 \text{ mol}$$

$$w = 0.1457 \text{ g}$$

Ni nitrate

Cr = $0.5 \times 0.03 = 0.015$ M

$$n = 0.00075 \text{ mol}$$

$$w = 0.300105 \text{ g}$$

Cr nitrate

Analytical Chemistry Calculation Questions And Answers

John Kenkel



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