

Binatone Fusion 2510 Answering Machine Manual

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 pK_b of the weak acid or base being titrated a solution pH as a function of the volume of 1.00 M NaOH added to 10.00 mL of
1.00 M solutions of weak acids with the indicated pK_a values
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determine the amount of substance in a solution of unknown concentration using various titrimetric methods theory titration

the word titration comes from the latin word titulus which means inscription or title

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web where the salt is the conjugate base of the acid the four parts of the titration curve are described below and you should look to the appropriate text section to see how they are treated pure acid 0 ml of base is added section 17 3 2 1

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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of hcl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0 100 m the ph of the solution is then $\text{ph} = -\log 0.100 = 1.000$ b titrant volume 12 50 ml

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web at the equivalence point in an acid base titration moles of base moles of acid and the solution only contains salt and water acid base titrations are monitored by the change of ph as titration progresses

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web expert answer 100 2 ratings the example is based upon henderson hasselbalch equation the equation gives relation between ph and pH_a of weak acid the equation is $\text{pH} = \text{pK}_a + \log \frac{[\text{salt}]}{[\text{acid}]}$ at equivalence point all a view the full answer transcribed image text

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web unlike strong acids or bases the shape of the titration curve for a weak acid or base depends on the pK_a or pK_b of the weak acid or base being titrated a solution ph as a function of the volume of 1 00 m naoh added to 10 00 ml of 1 00 m solutions of weak acids with the indicated pK_a values

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web titration curves source openstax chemistry atoms first 2e strong acid strong base titrations when titrating a strong acid such as hydrochloric acid with sodium hydroxide you are reacting the hcl directly with naoh $\text{hcl aq} + \text{naoh aq} \rightarrow \text{h}_2\text{O l} + \text{nacl aq}$ hcl a q naoh a q h 2 o l nacl a q

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titration curve is the plot of the pH value of a solution vs the amount of titrant added

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Volume of NaOH added (mL)	pH
4	4.6
6	4.6
5	9.5
9	5.9

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web background plot of the pH of a solution against the volume of titrant added is called a titration curve the pH can be measured directly with a pH meter while titrant is added from a buret for the acids used in this experiment the titrant will always be a 0.1 M solution of the strong base NaOH

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