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 $\text{H}_2\text{O}(\text{g}) + \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$
 $\text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) + \text{O}_2(\text{g}) \quad \Delta h = 283 \text{ kJ}$

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