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Chapter 5: Atom and Ion Movements in Materials

- 5-10 Atoms are found to move from one lattice position to another at the rate of 5×10^5 jumps/s at 400°C when the activation energy for their movement is $30,000 \text{ cal/mol}$. Calculate the jump rate at 750°C .

Solution:

$$\begin{aligned}\text{Rate} &= \frac{5 \times 10^5}{x} = \frac{c \exp[-30,000/(1.987)/(673)]}{c_v \exp[-30,000/(1.987)/(1023)]} = \exp(-22.434 + 14.759) \\ \frac{5 \times 10^5}{x} &= \exp(-7.675) = 4.64 \times 10^{-4} \\ x &= \frac{5 \times 10^5}{4.64 \times 10^{-4}} = 1.08 \times 10^9 \text{ jumps/s}\end{aligned}$$

- 5-11 The number of vacancies in a material is related to temperature by an Arrhenius equation. If the fraction of lattice points containing vacancies is 8×10^{-5} at 600°C , determine the fraction of lattice points containing vacancies at 1000°C .

Solution: $8 \times 10^{-5} = \exp[-Q/(1.987)/(873)]$ $Q = 16,364 \text{ cal/mol}$
 $f = n_v/n = \exp[-16,364/(1.987)/(1273)] = 0.00155$

- 5-12 The Arrhenius equation was originally developed for comparing rates of chemical reactions. Compare the rates of a chemical reaction at 20 and 100°C by calculating the ratio of the chemical reaction rates. Assume that the activation energy for liquids in which the chemical reaction is conducted is 10 kJ/mol and that the reaction is limited by diffusion.

Solution:

An Arrhenius relationship gives the rate of a process or reaction as a function of temperature. In general, for an Arrhenius relationship,

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**Donald R. Askeland, Frank
Haddleton, Phil Green, Howard
Robertson**



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The Science and Engineering of Materials Paul Porgess, Ian Brown, 2012-12-06 This solutions manual accompanies the SI edition of *The Science and Engineering of Materials* which emphasizes current materials testing procedures and selection and makes use of class tested examples and practice problems *Solutions Manual for the Science and Engineering of Materials* Donald R. Askeland, 1989 *The Science and Engineering of Materials* Donald R. Askeland, 1996-09-30 *The Science and Engineering of Materials* Donald R. Askeland, Frank Haddleton, Phil Green, Howard Robertson, 2013-11-11 The *Science and Engineering of Materials* Third Edition continues the general theme of the earlier editions in providing an understanding of the relationship between structure processing and properties of materials This text is intended for use by students of engineering rather than materials at first degree level who have completed prerequisites in chemistry physics and mathematics The author assumes these students will have had little or no exposure to engineering sciences such as statics dynamics and mechanics The material presented here admittedly cannot and should not be covered in a one semester course By selecting the appropriate topics however the instructor can emphasise metals provide a general overview of materials concentrate on mechanical behaviour or focus on physical properties Additionally the text provides the student with a useful reference for accompanying courses in manufacturing design or materials selection In an introductory survey texts such as this complex and comprehensive design problems cannot be realistically introduced because materials design and selection rely on many factors that come later in the student's curriculum To introduce the student to elements of design however more than 100 examples dealing with materials selection and design considerations are included in this edition

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