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Exercise 12

RESPIRATION: AEROBIC AND ANAEROBIC OXIDATION OF ORGANIC MOLECULES

This lab stresses the various reactions collectively termed respiration. Compare and contrast photosynthesis and respiration, and stress the significance of releasing energy in an efficient way that allows for its biological utilization. Also note that respiration is a defined sequence of complex reactions, and discuss the significance of aerobic vs. anaerobic respiration.

SUGGESTED ELEMENTS FOR AN INTRODUCTORY LECTURE

- Cellular respiration is the breaking of carbon bonds and the capture and storage of the released energy. This energy is stored primarily in the terminal phosphate bond of ATP and the high-energy electron carrier NADH.
- Reactions of glycolysis and the Krebs's cycle break carbon bonds and produce ATP and NADH as a cell's primary energy "currency."
- Fermentation uses NADH to reduce pyruvate and therefore regenerate NAD and recycle it to glycolysis reactions. Fermentation generates energy in the absence of oxygen.
- In the presence of oxygen NADH and FADH passes high-energy electrons to the Electron Transport System which generates ATP and H₂O.
- Aerobic organisms require oxygen to accept the electrons that have passed through the ETS and whose energy was sapped to phosphorylate ADP during the ETS.
- Anaerobes do not require oxygen. However, they can only produce small amounts of ATP (from glycolysis) because they cannot conduct the Krebs's Cycle and ETS in order to break all of the carbon bonds in a glucose molecule.
- Pyruvate is the final product of glycolysis. Fermentation is an anaerobic process that converts pyruvate to carbon dioxide and ethanol, or to lactic acid.

Biology Laboratory Manual 11th Edition Answers

Mary Burnham, Carol Hurd



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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the

copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web under the finalization of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away copper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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