

Are All Equilibria Created Equal?

Two bottles of soda are sitting in the refrigerator. One is almost full. The other is almost empty. Slowly, the sodas are going flat. This is because the dissolved carbon dioxide is coming out of solution and going into the air space above. Of course, if you keep the bottle closed tightly, the soda doesn't go flat as fast. If the gas can't escape, it soon reaches equilibrium with the solution below. But are all equilibria the same? Does the almost empty bottle retain as much carbonation as the almost full bottle? The soda with the larger air space has more room for the carbon dioxide gas to spread out. As a result the gas particles are under lower pressure and have a lower concentration. This makes it easier for them to come out of solution. When the almost empty bottle reaches equilibrium, it will have more carbon dioxide in the space above the liquid and less carbon dioxide dissolved than the almost full bottle has. It will be flat.

Equilibrium doesn't mean equal amounts of reactants and products. It means the reactants are turning into products at the same rate that the products are turning back to reactants. Many factors influence equilibrium. These include temperature, pressure, and concentration.



When sodas discuss solution equilibrium

Answer the questions below based on the reading above, and on your knowledge of chemistry.

1. A reversible reaction is at equilibrium. The forward reaction is exothermic.
 - a. Is the reverse reaction exothermic or endothermic? _____
 - b. Which has the higher activation energy, the forward reaction or the reverse? _____
 - c. What effect does raising the temperature have on the speed of reaction for the forward reaction? Why? _____

 - d. What effect does raising the temperature have on the speed of reaction for the reverse reaction? Why? _____

 - e. Will the rate of the forward reaction and the reverse reaction be affected to the same extent by an increase in temperature? Explain. _____

Are All Equilibria Created Equal Worksheet Answers

Joseph H. Noggle



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