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Al Sweigart



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water g sulfuric acid water h carbonate sulfuric acid copper ii carbonate water ii a $\text{ca oh}_2 \text{aq}$ $2\text{hno}_3 \text{aq}$ $\text{ca no}_3_2 \text{aq}$ $2\text{h}_2\text{o l}$

b zn s 2hcl aq $\text{zncl}_2 \text{aq}$ $\text{h}_2 \text{g}$ c 2koh aq

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chemical reaction $\text{nh}_3 \text{g}$ $\text{bf}_3 \text{g}$ $\text{nh}_3 \text{bf}_3 \text{s}$ show that the reaction illustrated by this equation is an acid base reaction if we

use the lewis definitions of an acid and a base see exercise 7

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acid base or both can be strong or weak electrolytes in aqueous solution and more

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the chemical definitions of an acid and a base in terms of their behavior in water write the formula of the salt formed when a given acid and base are combined

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web kw the product of the concentrations of the hydrogen ions and hydroxide ions acidic solution one which H^+ is greater than the OH^- basic solution one in which the H^+ is less than the OH^- alkaline solutions the H^+ of a basic solution is less than $1 \times 10^{-7} \text{ M}$

they are the same as basic solutions

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