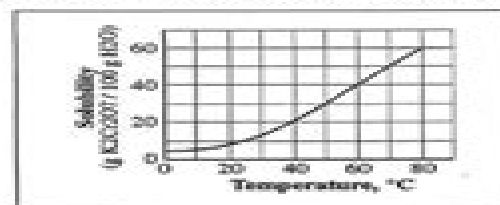


States of Matter / Solutions

- Two pure solid substances, **A** and **B**, each melt at 87.5 °C. An unknown pure solid substance **X** also melts at 87.5 °C. Melting points are obtained for mixtures of **X** with **A**, **X** with **B**, and **A** with **B**. The melting point of each mixture tested is found to be below 87.5 °C. Which conclusion is justified by these data?
 - X** is the same as **A** but different from **B**.
 - X**, **A**, and **B** are all the same substance.
 - X** is different from both **A** and **B**.
 - A** and **B** are the same, but **X** is different.
- A mixture of 0.5 mol of H_2 and 0.5 mol of SO_2 is introduced into a 10.0-L container at 25 °C. If the container has a pinhole leak, which describes the relationship between the partial pressures of the individual components in the container after 3 hours?
 - $\text{SO}_2 > \text{CH}_4 > \text{H}_2$
 - $\text{SO}_2 < \text{CH}_4 < \text{H}_2$
 - $\text{SO}_2 < \text{CH}_4 > \text{H}_2$
 - $\text{SO}_2 = \text{CH}_4 = \text{H}_2$
- Which microscopic representation best represents a solution?



- A mixture of 100 g of $\text{K}_2\text{Cr}_2\text{O}_7$ and 200 g of water is stirred at 60 °C until no more of the salt dissolves. The resulting solution is poured off, leaving the undissolved behind. The solution is now cooled to 20 °C. What mass $\text{K}_2\text{Cr}_2\text{O}_7$ crystallizes from the solution during the cooling?
 - 24 g
 - 31 g
 - 43 g
 - 62 g
- A student wants to prepare 250. mL of 0.10 M NaCl solution. Which procedure is the most appropriate?
 - Add 5.84 g of NaCl to 250. mL of H_2O
 - Add 1.46 g of NaCl to 250. mL of H_2O
 - Dissolve 5.84 g of NaCl in 50 mL of H_2O and dilute to 250. mL
 - Dissolve 1.46 g of NaCl in 50 mL of H_2O and dilute to 250. mL



Formula Molar Mass	
NaCl	58.4 g/mol

Your Answers:

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A red circular graphic with a gradient, appearing as a partial circle or a stylized 'C' shape, located to the right of the University of California, Santa Cruz text.

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