

Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium

Chapters 17 & 18

Essential Skills/ State Standards:

- 9-1. Describe how temperature and heat flow in terms of the motion of molecules (or atoms). (CSS: 7a)
 9-2. Know that chemical processes & phase changes can either release energy (exothermic) or absorb energy (endothermic). (CSS: 7b, 7c)
 9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat ($q = m \cdot c \cdot \Delta T$) & latent heat of phase change ($q = m \cdot \Delta H$). (CSS: 7d)
 9-4. Know the rate of reaction is the decrease in concentration of reactants or the increase in concentration of products with time & depends on such factors as: concentration, temperature, and pressure. (CSS: 8a, 8b)
 9-5. Know the definition and role of activation energy & catalyst in a chemical reaction. (CSS: 8d*, 8c)
 9-6. Know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure (CSS: 9a)
 9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (CSS: 9b)

Vocabulary

- | | | | |
|---|---|--|--|
| <ul style="list-style-type: none"> • Condensation • Freezing • Vaporization • Melting • Boiling • Sublimation | <ul style="list-style-type: none"> • Deposition • Heat • Temperature • Thermochemistry • Exothermic • Endothermic | <ul style="list-style-type: none"> • Calorimeter • Joule • Specific heat • $q = m \cdot c \cdot \Delta T$ • $q = m \cdot \Delta H_{\text{vaporization}}$ • $q = m \cdot \Delta H_{\text{fusion}}$ | <ul style="list-style-type: none"> • Heat of reaction • Le Chatelier's principle • Equilibrium • Activation Energy • Catalyst |
|---|---|--|--|

Book Assignments:

Topic	Read	Required Problems
1. Heat, Temperature & Specific Heat	Section 17-1 p. 511-524	Section 17-p. 524 #2, 5 & 6 Chapter 17 Review p. 548 #16 & 17
3. Chemical Equilibrium & Le Chatelier's Principle	Section 12-3 p. 372-375 Section 18-2 p. 562-568	Chapter 18 Review p. 586 #4, 6 & 7
4. Phase Changes Diagrams & Latent Heat of Phase Change	Section 12-3 p. 378-382	Section 12-3 p. 382 #4-8 Chapter 12 Review p. 389 #19, 28, 31 & 37
5. Collision Theory & Activation Energy	Section 17-3 p. 532-534	Section 17-3 p. 537 #2, 3 & 5
6. Reaction Rates & Catalysts	Section 17-4 p. 538-544	Section 17-4 p. 545 #2-4

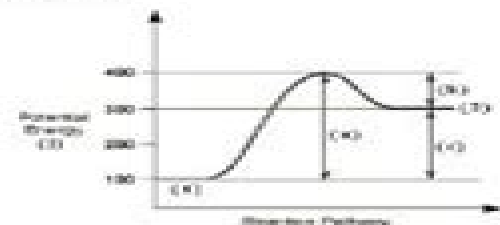
Study Guide (Use notes, worksheets, and labs in addition to this study guide)

- According to the collision theory, what 3 things must happen in order for two molecules to react?
- What is the importance of the activation energy in a chemical reaction?
- Using the collision theory, explain if the following would increase/ decrease the rate of reaction:

$$2 \text{HCl(aq)} + \text{Mg(s)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$$

a. Increase temp:	b. Increase HCl	c. Increase pressure	d. Remove MgCl_2	e. Add H_2
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- Identify the following terms with their correct number/letter from the graphs:

• Reactants-	• Products-
• Activation energy (number/letter & amount in Joules)-	
• Is this reaction exothermic or endothermic?	
- Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do).



Answers To Unit 8 Stoichiometry Study Guide

James E. Brady, John R. Holum



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Study Guide to Accompany Calculus for the Management, Life, and Social Sciences Clyde Metz, 1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences **Resources in Education**, 1997-05 **A**

Study Guide to Chemical Principles Wilbert Hutton, 1970 **Fundamentals of Chemistry, Study Guide** James E. Brady, John R. Holum, 1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs

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General Chemistry Study Guide Sixth Edition Darrell D. Ebbing, 1999 *Study Guide for General Chemistry and College Chemistry, Eighth Editions by Holtzclaw and Robinson* Norman E. Griswold, Henry Fuller Holtzclaw, 1988 **Laboratory**

Manual to Accompany Chemistry, [by] Stanley R. Radel, Marjorie H. Navidi Arthur D. Baker, 1990 **10 in One Study Package for CBSE Chemistry Class 12 with 5 Model Papers** Disha Experts, 2017-08-29 10 in ONE CBSE Study Package Chemistry class 12 with 5 Sample Papers is another innovative initiative from Disha Publication This book provides the excellent approach to Master the subject The book has 10 key ingredients that will help you achieve success 1 Chapter Utility Score 2 All India Board 2017 Solved Paper 3 Exhaustive theory based on the syllabus of NCERT books along with the concept maps for the bird's eye view of the chapter 4 NCERT Solutions NCERT Exercise Questions 5 VSA SA LA Questions Sufficient Practice Questions divided into VSA SA LA type Numericals are also included wherever required 6 Past Years Questions Past 10 year Questions of Board Exams are also included 7 HOTS Exemplar Value based Questions High Order Thinking Skill Based Moral Value Based and Selective NCERT Exemplar Questions included 8 Chapter Test A 15 marks test of 30 min to assess your preparation in each chapter 9 Important Formulae Terms and Definitions 10 Full syllabus Sample Papers 5 papers with detailed solutions designed exactly on the latest pattern of CBSE Board

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