



## Lab Data

*Verify your rate calculation.**Did you report your data to the correct number of significant figures?*

## Temperature

|                                                             | Trial 1                                   |
|-------------------------------------------------------------|-------------------------------------------|
| Concentration of iodide solution (mM)                       | 100.0                                     |
| Concentration of thiosulfate solution (mM)                  | 20.1                                      |
| Concentration of hydrogen peroxide solution (mM)            | 176.0                                     |
| Temperature of iodide solution (°C)                         | 4.6                                       |
| Volume of iodide solution ( $I^-$ ) used (mL)               | 10.0                                      |
| Volume of thiosulfate solution ( $S_2O_3^{2-}$ ) used (mL)  | 1.0                                       |
| Volume of hydrogen peroxide solution ( $H_2O_2$ ) used (mL) | 10.0                                      |
| Time (s)                                                    | 27.6                                      |
| Observations                                                | Solution turned blue in 27.6 seconds when |
| Initial concentration of iodide in reaction (mM)            | 47.6                                      |
| Initial concentration of thiosulfate in reaction (mM)       | 0.96                                      |
| Initial concentration of hydrogen peroxide in reaction (mM) | 83.8                                      |
| Initial Rate (mM/s)                                         |                                           |

# Answers To Iodine Clock Reaction Lab

**Ralph Petrucci, William  
Harwood, Geoffrey Herring**



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