

# 11

## ATOMIC SPECTROSCOPY

### Objectives

- To observe the characteristic emission frequencies of atoms and use these spectral lines to identify unknown sample ions.
- To calculate the energies of the emitted frequencies for hydrogen.
- To identify the electronic transitions and atomic energy levels from which these emissions arise.
- To calculate the value of the Rydberg constant.

### Background

According to modern quantum theory, the energy content of an atom is *quantized*, that is, the energy can have only certain discrete values. The allowed energy levels are different for different atoms and are characteristic of each kind of atom. Refer to your textbook for a more detailed discussion of this. Spectroscopy is the most important experimental tool for probing the allowed energy levels. In this experiment we will deal with *emission spectroscopy*, in which an atom in a high energy state undergoes a transition to a lower energy state. To rid itself of the excess energy, the atom emits light of a particular wavelength during the transition. The *spectrum* of an atom is the entire collection of all the wavelengths of light which can be emitted by the atom.

Quantum theory pictures light as being composed of tiny bundles of energy called *photons*. Different colors of light are produced by photons of different wavelengths. The energy of a photon,  $E$  (in Joules), is related to its wavelength,  $\lambda$  (in m), by

$$E = \frac{hc}{\lambda} \quad (1)$$

where  $h$  is Planck's constant,  $6.63 \times 10^{-34}$  J·s, and  $c$  is the speed of light in a vacuum,  $3.00 \times 10^8$  m/s.

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**S Nieto**



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