

- According to the Bronsted—Lowry definition, which chemical species can function *both* as an acid and as a base?
 - Cl^-
 - SO_4^{2-}
 - NH_4^+
 - HCO_3^-
 - H_3O^+
- In the reaction $\text{CN}^- + \text{H}_2\text{O} \rightleftharpoons \text{HCN} + \text{OH}^-$ which is an acid—base conjugate pair?
 - H_2O and HCN
 - H_2O and OH^-
 - CN^- and H_2O
 - HCN and OH^-
- Given that HX is a stronger Bronsted acid than HY in aqueous solution, which is true of a 1 M solution of NaX ?
 - It is less basic than a 1 M solution of NaY .
 - It is more basic than a 1 M solution of NaY .
 - It yields a neutral solution.
 - It is more concentrated than a 1 M solution of NaY .
- The pH of a 0.03 M HCl solution is
 - 1.5
 - 2.5
 - 3.5
 - 12.5
- The pH of a 1.0×10^{-2} M $\text{Ba}(\text{OH})_2$ solution at 25°C is
 - 2.7
 - 3.0
 - 11.0
 - 11.3
- Which is the strongest acid?
 - HClO
 - HClO_2
 - HClO_3
 - HClO_4
- Which salt reacts with water (hydrolyzes) to produce a basic solution?
 - $\text{NaC}_2\text{H}_3\text{O}_2$
 - NaNO_3
 - NH_4Cl
 - BaSO_4

- The *weakest* of the bases listed is

Acid	Conjugate Base	K_a (Ionization Constant of Acid)
HCl	Cl^-	100% ionized
HSO_4^-	SO_4^{2-}	1.2×10^{-2}
H_2S	HS^-	5.7×10^{-8}
HS^-	S^{2-}	1.2×10^{-13}

 - Cl^-
 - CN^-
 - HS^-
 - S^{2-}
 - SO_4^{2-}

- The oxide of which element will react with water to form the strongest acid?

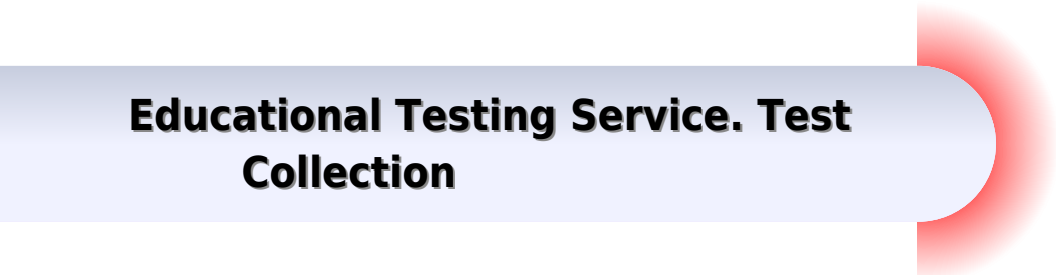
	Main Groups							
Period	I	II	III	IV	V	VI	VII	O
First								
Second	X	Y	Z	P		Q	S	U
Third	W					R	T	M

 - W
 - M
 - P
 - R
 - Z

- Which statement is a logical inference from the fact that a 0.10 M solution of potassium acetate $\text{KC}_2\text{H}_3\text{O}_2$ is less alkaline than a 0.10 M solution of potassium cyanide, KCN ?
 - Hydrocyanic acid is a weaker acid than acetic acid.
 - Hydrocyanic acid is less soluble in water than acetic acid.
 - Cyanides are less soluble than acetates.
 - Acetic acid is a weaker acid than hydrocyanic acid.
- In the titration of 50.0 mL of 0.100 M benzoic acid (a monoprotic acid) with 50.0 mL of 0.100 M NaOH , the properties of the solution at the equivalence point will correspond exactly to the properties of
 - a 0.100 M sodium solution.
 - a 0.0500 M sodium hydroxide solution.
 - a 0.0500 M benzoic acid solution.
 - a 0.0500 M sodium benzoate solution.

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