

WORKSHEET

Acid-base Titration I

Maximum score: 100 points

Your Name:

Partner:

Lab Section (Day & Time):

1. Attach your annotated plots and label the equivalence point and the half-equivalence points appropriately for the titration curves for HCl, KHP, and one trace for the unknown acid, and record the data from the plots. (30 points)

Acid Trial	pH at start of titration	pK _a or pH at half-equivalence point	pH at equivalence point	pH at end of plot
HCl	2.2	XXXXXXXXXXXX	5.3	11.75
KHP	4.6	5.1	10.5	11.70
Unknown	2.3	2.9	9.0	11.65

2. Complete the table using data from your lab notebook. Calculate the acid concentration using $M_a V_a = M_b V_b$. The unknown acid will require a correction for the dilution. (20 pts).

Acid Trial	Volume acid (mL)	NaOH Concentration (M)	Volume of base at Equivalence point (mL NaOH)	Known Acid Concentration (M)	Calculated Acid Concentration (M)
HCl	55	.1111	4.43	.1133	.0089
KHP	55	.1111	2.54	.0558	.0051
Diluted Unknowns Trial 1	55	.1111	9.005	XXXXXXXXXX	.0081
Diluted Unknowns Trial 2				XXXXXXXXXX	

Will not do second trial for unknown

Acid Base Titration Chemistry If8766 Answers

Sondra L. Spannuth



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Acid-base Titrations in Nonaqueous Solvents James Sherwood Fritz,1973 *Working with Chem Separate*
Wink,1999-12-01 **Acid-base Chemistry** Clifford W. Hand,Harry Lyon Blewitt,1986 Very Good No Highlights or Markup
all pages are intact **General Theory of Acid-base Titration** Pat Richard Pondy,1949 Titration in Non-aqueous
Media István Gyenes,1968 Aqueous Acid-base Equilibria and Titrations Robert De Levie,1999 This book will give students
a thorough grounding in pH and associated equilibria material absolutely fundamental to the understanding of many aspects
of chemistry It is in addition a fresh and modern approach to a topic all too often taught in an out moded way This book uses
new theoretical developments which have led to more generalized approaches to equilibrium problems these approaches are
often simpler than the approximations which they replace Acid base problems are readily addressed in terms of the proton
condition a convenient amalgam of the mass and charge constraints of the chemical system considered The graphical
approach of Bjerrum Hagg and Sillen is used to illustrate the orders of magnitude of the concentrations of the various species
involved in chemical equilibria Based on these concentrations the proton condition can usually be simplified often leading
directly to the value of the pH In the description of acid base titrations a general master equation is developed It provides a
continuous and complete description of the entire titration curve which can then be used for computer based comparison
with experimental data Graphical estimates of the steepness of titration curves are also developed from which the
practicality of a given titration can be anticipated Activity effects are described in detail including their effect on titration
curves The discussion emphasizes the distinction between equilibrium constants and electrometric pH measurements which
are subject to activity corrections and balance equations and spectroscopic pH measurements which are not Finally an entire
chapter is devoted to what the pH meter measures and to the experimental and theoretical uncertainties involved Ph
Scale And Acid-base Titrations/standardization Postma Roberts Hollenberg,James M. Postma,Julian Roberts,J. Leland
Hollenberg,2000-01-15 *Acid-base Indicators* Izaak Maurits Kolthoff,1937 The dissociation of strong and weak electrolytes
The properties of acid base indicators The colorimetric determination of hydrogen ion concentration *Lecture*
Demonstrations Employing Mechanical Reproductions of Acid-base Titration Curves and Photometric Titration Curves Sondra
L. Spannuth,1974 **Acid Base Titrations** Shah Purvesh,2015-08-04 In Acid Base Titrations the author discussed various
acid base titration It gives some basic concept of acid base titration The analysis of acid base titration discussed in this book
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