

SECTION 14.2

Acid-Base Theories

Teacher Notes and Answers

SECTION 2 ACID-BASE THEORIES

1. Brønsted-Lowry acids can donate protons to solutions other than water solutions.
2. Brønsted-Lowry bases accept donated protons. Arrhenius bases add hydroxide ions to aqueous solutions.
3. H_3O^+ , H_2PO_4^- , HPO_4^{2-} , PO_4^{3-}
4. The sodium ion is not a Lewis acid because it does not form a covalent bond. It accepts a single electron to form an ionic bond.
5. base; acid
6. Lewis acids

Review

1. Monoprotic acids can only donate one proton per molecule. Diprotic acids can donate two protons per molecule. Triprotic acids can donate three protons per molecule.
2. H_2CO_3 : proton donor, acidic; H_2O : proton acceptor, basic
3. AlCl_3 : electron pair acceptor, Lewis acid; H_2O : electron pair donor, Lewis base

- 4a. NaOH is an Arrhenius base because it increases the hydroxide ion concentration when it dissociates in water.
- 4b. H_2O is a Brønsted-Lowry base because it is a proton acceptor. It is also a Lewis base because it donates an electron pair to a hydrogen ion to form a hydronium ion.
- 4c. NH_3 is a Brønsted-Lowry base because it is a proton acceptor. It is also a Lewis base because it donates an electron pair to a hydrogen ion to form an ammonium ion.

Acid Base Theories Section Review Answers

Marcel A. Müller



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