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#### Chapter 5: Atom and Ion Movements in Materials

5-10 Atoms are found to move from one lattice position to another at the rate of 5 × 10<sup>3</sup> jumps/s at 400°C when the activation energy for their movement is 30,000 cal/mol. Calculate the jump rate at 750°C.

#### Solution:

Rate = 
$$\frac{5 \times 10^3}{x} = \frac{c \exp[-30,000/(1.987)/(673)]}{c \exp[-30,000/(1.987)/(1023)]} = \exp(-22.434 + 14.759)$$
  
 $\frac{5 \times 10^3}{x} = \exp(-7.675) = 4.64 \times 10^{-4}$   
 $x = \frac{5 \times 10^3}{4.64 \times 10^{-4}} = 1.08 \times 10^9 \text{ jumps/s}$ 

5-11 The number of vacancies in a material is related to temperature by an Arrhenius equation. If the fraction of lattice points containing vacancies is 8 × 10<sup>-6</sup> at 600°C, determine the fraction of lattice points containing vacancies at 1000°C.

Solution: 
$$8 \times 10^{-5} = \exp[-Q/(1.987)/(873)]$$
  $Q = 16,364 \text{ cal/mol}$   $f = n_n/n = \exp[-16,364/(1.987)/(1273)] = 0.00155$ 

5-12 The Arrhenius equation was originally developed for comparing rates of chemical reactions. Compare the rates of a chemical reaction at 20 and 100°C by calculating the ratio of the chemical reaction rates. Assume that the activation energy for liquids in which the chemical reaction is conducted is 10 kJ/mol and that the reaction is limited by diffusion.

#### Solution:

An Arrhenius relationship gives the rate of a process or reaction as a function of temperature. In general, for an Arrhenius relationship,

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