

Acids & Bases Study Guide

6. Water is considered neutral. Describe what this means in terms of $[H^+]$ & $[OH^-]$ concentrations.

8. Complete the following:

$$pH = -\log[H^+]$$

$$K_w = [H^+][OH^-]$$

7. Fill in the blank:

9. What is the $[OH^-]$ of a solution with a pH of 10?

10. What is the $[H^+]$ of a solution with a pH of 4?

11. What is the pH of a solution with a $[H^+]$ of 1×10^{-5} M?

12. What is the pH of a solution with a $[OH^-]$ of 1×10^{-4} M?

13. What is the pH of a solution with a $[H^+]$ of 1×10^{-12} M?

14. What is the pH of a solution with a $[OH^-]$ of 1×10^{-10} M?

15. What is the pH of a solution with a $[H^+]$ of 1×10^{-11} M?

16. What is the pH of a solution with a $[OH^-]$ of 1×10^{-13} M?

Name: _____ Period: _____ Date: _____

Acids & Bases Study Guide

1. Fill in the blanks or complete and contrast the properties of Acids & Bases.

Acids	Bases
• Corrosive _____ • Protons _____ • Tastes _____ • Turn litmus paper _____ • High $[H^+]$ _____	• Corrosive _____ • Protons _____ • Tastes _____ • Turn litmus paper _____ • High $[OH^-]$ _____

2. Fill in the chart for the following acids & bases.

Acid or Base?	Strong or Weak?	Formula	Name
			hydrochloric acid
			barium hydroxide
		$HClO_4$	
		$NaOH$	
		H_2SO_4	hydrochloric acid

3. Predict the products for the following double replacement reactions involving acids & bases. Identify which component is the Acid, Base, Conjugate Acid (CA), and Conjugate Base (CB) in the formula.



4. Label the pH Scale from 0 (Most Basic) to 14 (Most Acidic). (Strong Acid, Weak Acid, Strong Base, Weak Base)



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O. Ray Whittington

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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