

Pre-Laboratory Assignment: Acid-Base Titrations

Name _____ Section _____ Date _____

1. Find in your textbook the dilution formula and calculate the amount of 6M NaOH required to prepare 500 ml of 0.25 M NaOH. 2 decimals

$$M_1 = 6M \text{ NaOH}$$

$$V_1 = ?$$

$$M_2 = 0.25M$$

$$V_2 = 500mL$$

$$V_1 = \frac{M_2 V_2}{M_1}$$

$$V_1 = \frac{0.25 \times 500}{6.0} = 20.83mL$$

2. Repeat the calculation for the amount of 6 M HCl required to prepare 300 mL of 0.35 M HCl. 2 decimals

$$M_1 = 6M \text{ HCl}$$

$$V_1 = ?$$

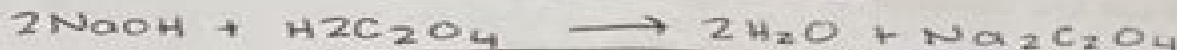
$$V_2 = 300mL$$

$$M_2 = 0.35M$$

$$V_1 = \frac{300 \times 0.35}{6.0} = 17.50$$

3. Calculate the mass of oxalic acid dihydrate ($H_2C_2O_4 \cdot 2H_2O$) required to react with 15. mL of 0.25 M sodium hydroxide solution. (This is the calculation, needed in Part III, of the sample size of your primary standard for the standardization of NaOH.) 1 decimal

$$0.015 \times 0.25 = 0.00375 \text{ mol NaOH}$$



$$0.00375 \text{ mol NaOH} \times \frac{1}{2} = 0.001875 \text{ mol } H_2C_2O_4$$

$$\text{mol } H_2C_2O_4 \cdot 2H_2O = 0.001875$$

$$0.001875 \times 126.07 = 0.236g$$

$$\text{Mass } H_2C_2O_4 \cdot 2H_2O = 0.2g$$

Acid Base Titration Pre Lab Answers Uic

**Donald J. Wink, Sharon Fetzer-
Gislason, Julie Ellefson Kuehn**



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Working with Chemistry Donald J. Wink, Sharon Fetzner-Gislason, Julie Ellefson Kuehn, 2004-02-20 With this modular laboratory program students build skills using important chemical concepts and techniques to the point where they are able to design a solution to a scenario drawn from a professional environment The scenarios are drawn from the lives of people who work with chemistry every day ranging from field ecologists to chemical engineers and include many health professionals as well The Effect of Adding Guided-inquiry to Laboratory Activities in an Acid Base Unit in a High School Chemistry Classroom Kendon Douglas Smith, 2010 Chemistry Lab Investigations Saurya Singh, 2021-01-17 The book provides coverage of the essential lab topics of the AP and IB Chemistry courses Each lab investigation is well structured with an introduction lab concepts procedure execution results analysis and conclusion The key lab investigations in the book are Identifying the types of solids and the forces in action by physical properties Investigating the mole ratio in a chemical reaction Separating the solutes from a mixture using chromatography Finding out the amount of phosphate in plant food Simulating and analyzing the bond polarity partial charges and electrostatic forces using electronegativity Investigating the reversible reaction and applied Le Chatelier's principle Performing acid base titration to observe pH curve and investigating the properties of the buffer solution Finding oxidation states using redox titration Constructing a galvanic cell and determining the cell voltage

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