

FINAL EXAM STUDY GUIDE

▼ CHAPTER 1: INTRO TO ACCOUNTING

▼ GAAP

- generally accepted accounting principles
- standards that are generally accepted and universally practiced; indicate how to report economic events
- financial statements: balance sheet, income statement, statement of SE, statement of cash flows, note disclosure
- standard setting bodies: FASB, SEC, IASB

▼ basic accounting equation

- assets = liabilities + stockholder's equity
- assets → what a business owns, helps a company make money, provide future services or benefits (cash, supplies, equipment, lands, buildings, computers, vehicles, AR)
- liabilities → debts and obligations (what a company owes (AP, NP, SWP, sales and real estate taxes payable, etc.))
- SE → ownership claims on total assets, net worth (common stock, retained earnings, revenues, expenses, dividends); what you own minus what you owe

▼ income statement

- reports profitability of the company's operations over a specific period of time; shows net income or net loss

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Wiley



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web $m \times c \times \Delta t$ equation 1 where m is the total mass of the solution solute plus solvent c is the specific heat of the solution and Δt is the observed temperature change the specific heat of the solution is generally assumed to be the same as that of

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web the heat of reaction ΔH_{soln} is written after the products in units of kJ mol rxn for example sodium hydroxide dissolves exothermically releasing 44.2 kilojoules per mole dissolved this process is represented as $\text{NaOH (s)} \rightarrow \text{NaOH (aq)}$ $\Delta H_{\text{soln}} = -44.2 \text{ kJ/mol rxn}$ write equations to similarly represent the dissolving process for

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web hand warmers are usually placed inside the gloves to keep your hands warm the heat keeping your hands warm is formed as a result of an exothermic reaction which occurs in a small packet hand warmers have 5 main ingredients in them salt carbon polypropylene vermiculite water and iron

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web background the hand warmers are filled with sodium acetate click the metal disc and you will notice the liquid changes state and becomes a solid in the process heat is released the sodium acetate is a supersaturated solution

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web oct 9 2021 sodium acetate coin or other small piece of metal plastic bags seal the coin inside a small bag place this bag inside a larger bag containing supersaturated sodium acetate get this by dissolving as much sodium acetate as possible in hot water this chemical hand warmer is reusable

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web the goal of my experiment and research is to evaluate the various number of compounds to be used as hand warmers and locate the ones that can raise the temperature to 20 degrees celsius in 100 g water and costs as little as possible

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web in this lab you are challenged to use chemistry to determine which ionic compound makes the best hand warmer in this experiment you will collect data that will allow you to calculate the heat of solution ΔH_{soln} or the amount of energy released or absorbed in kJ/mol of the substance being dissolved

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web mar 4 2014 contains water the hand warmer is struck in a manner that ruptures the inner pouch releasing the ionic salt into the water of the outer pouch the salt dissolves and the water warms this type of hand warmer tends to produce a more vigorous heat than the dry powder type of hand warmer but does not produce heat for quite as long

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initiates the bow bending from the waist to an angle of between 30 and 45 degrees from vertical a less accentuated bow is
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