

Задача 11.
Слабые и сред. кислоты

① Дано:

$$C_{\text{HNO}_2} = 0,01 \text{ м}$$

$$K_a = 7,9 \cdot 10^{-4}$$

$$pH = ?$$

$$pH = -\lg [H^+] = -\lg 2,8 \cdot 10^{-5} = 4,55$$

Решение:

Для слабой кислоты:

$$[H^+] = \sqrt{K_a \cdot C_{\text{HNO}_2}}$$

$$= \sqrt{7,9 \cdot 10^{-4} \cdot 10^{-2}} = 2,8 \cdot 10^{-5} \text{ моль/л}$$

$$= 4,55$$

② Дано:

$$K_{\text{PrOH}_2} = 6,3 \cdot 10^{-5}$$

$$pH = ?$$

Решение:

$$C_{\text{PrOH}_2} = C_{\text{H}^+} = C_{\text{OH}^-} = 2,5 \cdot 10^{-3}$$

$$K_{\text{Pr}} = [C_{\text{H}^+}][C_{\text{OH}^-}] = 2,5 \cdot 10^{-3} \cdot 2,5 \cdot 10^{-3} = 6,25 \cdot 10^{-6}$$

$$= 6,25 \cdot 10^{-6}$$

3 - гидроксида моль/л

$$S = \sqrt[3]{\frac{K_{\text{Pr}}}{4}} = \sqrt[3]{\frac{6,25 \cdot 10^{-6}}{4}} = 1,16 \cdot 10^{-2} \text{ моль/л}$$

$$[C_{\text{H}^+}] = \frac{S}{2} = \frac{1,16 \cdot 10^{-2}}{2} = 5,8 \cdot 10^{-3} \text{ моль/л}$$

$$[C_{\text{H}^+}] = \frac{10^{-14}}{[C_{\text{OH}^-}]} = \frac{10^{-14}}{2,5 \cdot 10^{-3}} = 4 \cdot 10^{-12}$$

$$pH = -\lg [H^+] = -\lg 4 \cdot 10^{-12} = 11,4$$

③ Дано:

$$V_{\text{HNO}_2} = 10 \text{ мл}$$

$$C_{\text{HNO}_2} = 0,1 \text{ м}$$

$$V_{\text{PrOH}_2} = 10 \text{ мл}$$

$$C_{\text{PrOH}_2} = 0,1 \text{ м}$$

$$K_{\text{PrOH}_2} = 6,3 \cdot 10^{-5}$$

$$K_{\text{PrOH}_2} = 6,3 \cdot 10^{-5}$$

Решение:

полн. смешивание:

$$C_{\text{HNO}_2} = \frac{0,1 \cdot 10}{20} = 0,05 \text{ моль/л}$$

$$C_{\text{PrOH}_2} = \frac{0,1 \cdot 10}{20} = 0,05 \text{ моль/л}$$

или гидролизирован. комплекс

$$C_{\text{H}^+} = C_{\text{HNO}_2} = 0,05 \text{ моль/л}$$

$$C_{\text{PrOH}_2} = C_{\text{PrOH}_2} = 0,05 \text{ моль/л}$$

$$[C_{\text{H}^+}][C_{\text{PrOH}_2}] = 6,3 \cdot 10^{-5} = 10^{-14}$$

$$PrOH_2 \rightleftharpoons PrOH_2^+ + OH^-$$

$$K_{\text{PrOH}_2} = 6,3 \cdot 10^{-5}$$

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Albert A Gayle



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Table of Contents Apex Learning Chemistry Semester 2 Answers

1. Understanding the eBook Apex Learning Chemistry Semester 2 Answers
 - The Rise of Digital Reading Apex Learning Chemistry Semester 2 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Apex Learning Chemistry Semester 2 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Apex Learning Chemistry Semester 2 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Apex Learning Chemistry Semester 2 Answers
 - Personalized Recommendations
 - Apex Learning Chemistry Semester 2 Answers User Reviews and Ratings
 - Apex Learning Chemistry Semester 2 Answers and Bestseller Lists
5. Accessing Apex Learning Chemistry Semester 2 Answers Free and Paid eBooks
 - Apex Learning Chemistry Semester 2 Answers Public Domain eBooks
 - Apex Learning Chemistry Semester 2 Answers eBook Subscription Services

- Apex Learning Chemistry Semester 2 Answers Budget-Friendly Options
- 6. Navigating Apex Learning Chemistry Semester 2 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Apex Learning Chemistry Semester 2 Answers Compatibility with Devices
 - Apex Learning Chemistry Semester 2 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Apex Learning Chemistry Semester 2 Answers
 - Highlighting and Note-Taking Apex Learning Chemistry Semester 2 Answers
 - Interactive Elements Apex Learning Chemistry Semester 2 Answers
- 8. Staying Engaged with Apex Learning Chemistry Semester 2 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Apex Learning Chemistry Semester 2 Answers
- 9. Balancing eBooks and Physical Books Apex Learning Chemistry Semester 2 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Apex Learning Chemistry Semester 2 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Apex Learning Chemistry Semester 2 Answers
 - Setting Reading Goals Apex Learning Chemistry Semester 2 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Apex Learning Chemistry Semester 2 Answers
 - Fact-Checking eBook Content of Apex Learning Chemistry Semester 2 Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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