

RADIOACTIVE DECAY

Radioactive Decay of Candium 150

Introduction: Some isotopes spontaneously emit particles or energy for stable or unstable elements. The unstable radioactive decays. This decay occurs naturally and is not affected by chemical change, temperature, or pressure.

Radioactive decay takes place of a random rate. Although you cannot predict just one given atom will decay, you can predict that all billions of atoms will be small pieces of a radioactive element, a given number will decay at a predictable and predictable rate.

Testing of radioactive elements in rocks have determined the absolute age of the rock by comparing the percentage of original element present today to the percentage of the element present today, the age of which can be calculated. The ratio of the two atoms is a direct function of time because when the rock was formed, it must have contained a certain number of daughter atoms.

Procedure:

- Candium 150
- Half-life
- Radioactive decay
- Radioactive dating

4. Count and record the number of undecayed "undecayed" Candium atoms (M) and the remaining.

a. Record in the data table.

Repeat steps 2, 3 and 4 until all the radioactive "undecayed" (Parent M) atoms decay or 10 minutes in the case of radioactive decay.

Half-life (# of trials)	"Parent" atoms (# of "undecayed" radioactive Candium atoms remaining after the "M" side up)	"Daughter" atoms (# of "decayed" atoms)
0	50	0
1		
2		
3		
4		
5		
6		
7		
8		
9		
10		

Graph: Use the graph on the following page to represent the data collected throughout the procedure. The "parent" atoms (green) and "daughter" atoms (red) should be plotted on your graph.

SPECTACULAR SCIENCE

STUDENT LAB

Activity 1 Radioactive Decay Lab Student Guide

Robert R. Redfield



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