

ABBREVIATIONS AND SYMBOLS			
amount of substance	n	joule	J
atmosphere	atm	kelvin	K
boiling pt. elev. const.	K_b	liter	L
Celsius temperature	$^{\circ}\text{C}$	molar	M
energy of activation	E_a	molar	M
enthalpy	H	mole	mol
entropy	S	pressure	P
equilibrium constant	K_{eq}	second	s
free energy	G	specific heat capacity	C_p
freezing pt. depr. const.	K_f	temperature	T
gram	g	volt	V
hour	h	volume	V

CONSTANTS	
Planck's constant $h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$	Gas Constant $R = 0.0821 \text{ (L}\cdot\text{atm)/(mol}\cdot\text{K)}$ $R = 8.314 \text{ J/(mol}\cdot\text{K)}$
Speed of light $c = 2.998 \times 10^8 \text{ m}\cdot\text{s}^{-1}$	Water Dissociation Constant $K_w = 1.0 \times 10^{-14} \text{ at } 25^{\circ}\text{C}$
Rydberg constant $R_H = 2.18 \times 10^{-18} \text{ J}$	Faraday Constant $F = 96470 \text{ C/mol e}^{-}$ $F = 96470 \text{ A}\cdot\text{s/mol e}^{-}$ $F = 96470 \text{ J/V}\cdot\text{mol e}^{-}$
Avogadro Constant $N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$	
Measures of pressure 1 atm = 760 mmHg	

EQUATIONS			
$E = h\nu \quad E = \frac{hc}{\lambda}$	$\frac{\text{rate}_1}{\text{rate}_2} = \sqrt{\frac{\text{molar mass}_1}{\text{molar mass}_2}}$	Integrated Rate Laws zero: $[A] = [A]_0 - kt$	$\text{pH} = \text{p}K_a + \log\left(\frac{[A^-]}{[HA]}\right)$
$\Delta E = R_{\infty} \left(\frac{1}{n_1^2} - \frac{1}{n_2^2} \right)$	$\Delta T_1 = K_1 m \quad \Delta T_2 = i K_1 m$ $\Delta T_b = K_b m \quad \Delta T_b = i K_b m$ $\Pi = MRT \quad \Pi = i MRT$	second: $\frac{1}{[A]} = \frac{1}{[A]_0} + kt$	$\Delta G^{\circ} = \Delta H^{\circ} - T\Delta S^{\circ}$
$q = m c \Delta T$ $PV = nRT$	$\ln\left(\frac{k_2}{k_1}\right) = \frac{E_a}{R} \left(\frac{1}{T_1} - \frac{1}{T_2} \right)$	first: $\ln[A] = \ln[A]_0 - kt$	$\Delta G^{\circ} = -RT \ln K \quad \Delta G = \Delta G^{\circ} + RT \ln Q$ $E^{\circ} = \frac{RT}{nF} \ln K \quad E = E^{\circ} - \frac{RT}{nF} \ln Q$

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RJ Shavelson



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