

ACIDS, BASES AND SALTS

Definition of an acid	What causes the acidic properties of acids?
An acid is a compound which when dissolved in water produces hydrogen ions as the only positively charged ions	The hydrogen ions (H^+) cause the acidic properties and these are formed in the presence of water.

Another term to refer to an acid	Why an acid is also called a proton donor?
An acid is called a <u>proton donor</u>	It's because an acid provides protons or hydrogen ions (H^+) to other substances during the reaction.

Substance to which an acid provides protons	A base is then called	Why the base is called a proton acceptor?	Equation for the reaction between acid and base
BASE	PROTON ACCEPTOR	Because the base accepts hydrogen ions from acids	$H^+_{(aq)} + OH^-_{(aq)} \longrightarrow H_2O_{(l)}$

Common laboratory acids	These three common laboratory acids are also called	Why these acids are also called mineral acids?
Hydrochloric acid	HCl	They are derived from mineral salts i.e. chlorides for HCl, sulphates for H_2SO_4 and nitrates for HNO_3
Sulphuric acid	H_2SO_4	
Nitric acid	HNO_3	

Other mineral acids known	Mineral salts from which the acid is derived	Organic acids known	Naturally occurring acids
Sulphurous acid	H_2SO_3	Ethanoic acid (CH_3COOH)	CITRIC ACID from lemons
Carbonic acid	H_2CO_3		TARTARIC ACID from grapes
Phosphoric acid	H_3PO_4	Methanoic acid ($HCOOH$)	ACETIC ACID from vinegar
Nitrous acid	HNO_2		LACTIC ACID from sour milk
			Hydrochloric acid from digestive juices

Whenever an acid is dissolved in water, it produces	Term given to the number of hydrogen ions produced by one molecule of an acid	Definition of basicity of an acid
HYDROGEN IONS	BASICITY OF AN ACID	BASICITY of an acid is the <u>number of hydrogen ions</u> produced by <u>one molecule of an acid in aqueous solution</u> .

Acids And Bases Study Guide

RC Schank



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