

## Acids and Bases Worksheet

1. Acids and bases share some properties but are different in others. Use the samples at the side of the class to complete the chart. If both acids and bases share the property, place it in both columns.

| Property                         | Acids               | Bases    |
|----------------------------------|---------------------|----------|
| Conductivity in Aqueous Solution | very                | very     |
| Taste                            | Sour                | Bitter   |
| Feel                             | —                   | Slippery |
| Reaction with Metal              | produces $H_2$ gas  | —        |
| Reaction with Carbonate          | produces $CO_2$ gas | —        |
| Colour on Litmus Paper           | Red                 | Blue     |
| Colour with Universal Indicator  | Red                 | Purple   |
| Colour with phenolphthalein      | Colourless          | pink     |

2. Indicate if each compound is an acid or base and name it.

|            | Acid/Base? | Name                        |
|------------|------------|-----------------------------|
| $HNO_3$    | <u>A</u>   | <u>nitric acid</u>          |
| $Mg(OH)_2$ | <u>B</u>   | <u>magnesium hydroxide</u>  |
| $H_3PO_4$  | <u>A</u>   | <u>phosphoric acid</u>      |
| $H_2CO_3$  | <u>A</u>   | <u>carbonic acid</u>        |
| $HF$       | <u>A</u>   | <u>hydrofluoric acid</u>    |
| $Fe(OH)_3$ | <u>B</u>   | <u>iron (iii) hydroxide</u> |

3. Write the formula for each of the following acids/bases:

|                    |                              |                   |                              |
|--------------------|------------------------------|-------------------|------------------------------|
| hydrochloric acid  | <u><math>HCl</math></u>      | hydrochloric acid | <u><math>HCl</math></u>      |
| aluminum hydroxide | <u><math>Al(OH)_3</math></u> | sodium hydroxide  | <u><math>NaOH</math></u>     |
| lithium hydroxide  | <u><math>LiOH</math></u>     | calcium hydroxide | <u><math>Ca(OH)_2</math></u> |
| sulfuric acid      | <u><math>H_2SO_4</math></u>  | sulfurous acid    | <u><math>H_2SO_3</math></u>  |

# Acids And Bases Guide Answers

**S Ben Porath**



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## **Table of Contents Acids And Bases Guide Answers**

1. Understanding the eBook Acids And Bases Guide Answers
  - The Rise of Digital Reading Acids And Bases Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Acids And Bases Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Acids And Bases Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Acids And Bases Guide Answers

- Personalized Recommendations
- Acids And Bases Guide Answers User Reviews and Ratings
- Acids And Bases Guide Answers and Bestseller Lists
- 5. Accessing Acids And Bases Guide Answers Free and Paid eBooks
  - Acids And Bases Guide Answers Public Domain eBooks
  - Acids And Bases Guide Answers eBook Subscription Services
  - Acids And Bases Guide Answers Budget-Friendly Options
- 6. Navigating Acids And Bases Guide Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Acids And Bases Guide Answers Compatibility with Devices
  - Acids And Bases Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Acids And Bases Guide Answers
  - Highlighting and Note-Taking Acids And Bases Guide Answers
  - Interactive Elements Acids And Bases Guide Answers
- 8. Staying Engaged with Acids And Bases Guide Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Acids And Bases Guide Answers
- 9. Balancing eBooks and Physical Books Acids And Bases Guide Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Acids And Bases Guide Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Acids And Bases Guide Answers
  - Setting Reading Goals Acids And Bases Guide Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Acids And Bases Guide Answers

- Fact-Checking eBook Content of Acids And Bases Guide Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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