

Acid-Base Equilibrium Homework Part II

Part A

1. Calculate the pH and pOH of a solution containing 7.30 grams of HCl in 1.0 L of solution.

$$7.30 \text{ g} \times \frac{1 \text{ mol}}{36.46 \text{ g}} = 0.200 \text{ mol} / 1.0 \text{ L} = 0.20 \text{ M HCl}$$

Because HCl is a strong acid, it completely dissociates, so $[\text{HCl}] = [\text{H}_3\text{O}^+] = 0.20 \text{ M}$

$$\text{pH} = -\log[0.20] = 0.70$$

$$\text{pOH} = 14 - \text{pH} = 13.30$$

2. Calculate the pH of a 0.10 M nitric acid (HNO_3) solution.

Because HNO_3 is a strong acid, the acid completely dissociates and the concentration of the acid is equal to the concentration of H^+ at equilibrium.

$$[\text{H}_3\text{O}^+] = 0.10 \text{ M}$$

$$\text{pH} = -\log(0.10) = 1$$

3. Calculate the pH of a $5.0 \times 10^{-3} \text{ M}$ solution of sodium hydroxide.

Sodium hydroxide is a strong base, which means it completely dissociates, so $[\text{NaOH}] = [\text{OH}^-]$.

$$[\text{OH}^-] = 5.0 \times 10^{-3}$$

$$\text{pOH} = -\log(5.0 \times 10^{-3}) = 1.3$$

$$\text{pH} = 14 - 1.3 = 12.7 = 13 \text{ (2 sig figs)}$$

Part B

4. Decide whether the statements are true about strong acids, weak acids, both, or neither.

- a. Always dissociates leaving behind at least some hydronium ion **both**
- b. The initial concentration of the acid is equivalent to the final concentration of hydronium ions **strong only**
- c. The pH is easily calculated **strong only**
- d. The pH must be calculated by the ICE method **weak only**
- e. The acid completely dissociates in water **strong only**
- f. The equilibrium concentration of undissociated acid is approximately 0 **strong only**
- g. K_a is very small **weak only**

5. Write the K_a expression for the following acids:

- a. $\text{HC}_2\text{H}_3\text{O}_2$



$$K_a = \frac{[\text{C}_2\text{H}_3\text{O}_2^-][\text{H}_3\text{O}^+]}{[\text{HC}_2\text{H}_3\text{O}_2]}$$

- b. H_2CO_3



$$K_a = \frac{[\text{HCO}_3^-][\text{H}_3\text{O}^+]}{[\text{H}_2\text{CO}_3]}$$

Acid Base Homework Answers

Nurse Sarah



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Acids & Bases Carol Baldwin, 2005-10-11 Describes the chemical principles of acids and bases suggesting experiments to show how they can be neutralized and put to use *Homework Helpers: Chemistry* Greg Curran, 2025-09-12 Homework Helpers Chemistry is a user friendly review book that will make every student or parent trying to help their child feel like he or she has a private Chemistry tutor Concepts are explained in clear easy to understand language and problems are worked out with step by step methods that are easy to follow Each lesson comes with numerous review questions and answer keynotes that explain each correct answer and why it is correct This book covers all of the topics in a typical one year Chemistry curriculum including A systematic approach to problem solving conversions and the use of units Naming compounds writing formulas and balancing chemical equations Gas laws chemical kinetics acids and bases electrochemistry and more While Homework Helpers Chemistry is an excellent review for any standardized Chemistry test including the SAT II its real value is in providing support and guidance during the year's entire course of study **Organic Chemistry** Michael B. Smith, 2016-03-09 Based on the premise that many if not most reactions in organic chemistry can be explained by variations of fundamental acid base concepts Organic Chemistry An Acid Base Approach provides a framework for understanding the subject that goes beyond mere memorization Using several techniques to develop a relational understanding it helps students fully grasp the essential concepts at the root of organic chemistry This new edition was rewritten largely with the feedback of students in mind and is also based on the author's classroom experiences using the first edition Highlights of the Second Edition Include Reorganized chapters that improve the presentation of material Coverage of new topics such as green chemistry Adding photographs to the lectures to illustrate and emphasize important concepts A downloadable solutions manual The second edition of Organic Chemistry An Acid Base Approach constitutes a significant improvement upon a unique introductory technique to organic chemistry The reactions and mechanisms it covers are the most fundamental concepts in organic chemistry that are applied to industry biological chemistry biochemistry molecular biology and pharmacy Using an illustrated conceptual approach rather than presenting sets of principles and theories to memorize it gives students a more concrete understanding of the material **Acids and Bases** Carol Baldwin, 2004 Examines the properties of acids and bases where these compounds are found and how they interact with chemicals *ABG Interpretation: Notes, Mnemonics, & Workbook* by Nurse Sarah Nurse Sarah, ABG Interpretation Notes Mnemonics Workbook by Nurse Sarah This eBook contains Nurse Sarah's arterial blood gas notes and memory tricks to help you learn and retain important ABG concepts for nursing school nursing exams and while working as a nurse Nurse Sarah has condensed and illustrated her YouTube lectures into easy to read digital study notes that you can download INSTANTLY and access on your devices for quick review After your purchase you will be provided with information to download the eBook Please NOTE This eBook is in PDF format 22 MB This eBook contains a total of 64 pages with beautiful colors and

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