

Pre-Laboratory Assignment: Acid-Base Titrations

Name _____ Section _____ Date _____

1. Find in your textbook the dilution formula and calculate the amount of 6M NaOH required to prepare 500 ml of 0.25 M NaOH. 2 decimals

$$M_1 = 6M \text{ NaOH}$$

$$V_1 = ?$$

$$M_2 = 0.25M$$

$$V_2 = 500mL$$

$$V_1 = \frac{M_2 V_2}{M_1}$$

$$V_1 = \frac{0.25 \times 500}{6.0} = 20.83mL$$

2. Repeat the calculation for the amount of 6 M HCl required to prepare 300 mL of 0.35 M HCl. 2 decimals

$$M_1 = 6M \text{ HCl}$$

$$V_1 = ?$$

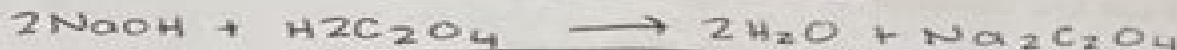
$$V_2 = 300mL$$

$$M_2 = 0.35M$$

$$V_1 = \frac{300 \times 0.35}{6.0} = 17.50$$

3. Calculate the mass of oxalic acid dihydrate ($H_2C_2O_4 \cdot 2H_2O$) required to react with 15. mL of 0.25 M sodium hydroxide solution. (This is the calculation, needed in Part III, of the sample size of your primary standard for the standardization of NaOH.) 1 decimal

$$0.015 \times 0.25 = 0.00375 \text{ mol NaOH}$$



$$0.00375 \text{ mol NaOH} \times \frac{1}{2} = 0.001875 \text{ mol } H_2C_2O_4$$

$$\text{mol } H_2C_2O_4 \cdot 2H_2O = 0.001875$$

$$0.001875 \times 126.07 = 0.236g$$

$$\text{Mass } H_2C_2O_4 \cdot 2H_2O = 0.2g$$

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