

# Solutions, acids, and bases

## 1. VOCABULARY REVIEW OF SOLUTIONS

**Solute** - what is **dissolved** in the solvent to form a solution.

**Solvent** - the substance that **dissolves** the solute, ordinarily a liquid when the solute and the solvent come together to form a **homogeneous mixture**.

**Miscible** - two mix in all proportions (want to be mixed) - liquids will pour each other

**Immiscible** - don't mix in any proportions (want to be mixed)

**Examples of miscible/immiscible** - ethanol/water/methanol, etc. and hexane/octane

**Solute** - aspect of being **dissolved** in liquid

**Insoluble** - a substance (want to be a solid) that **will not dissolve** in a solvent - an example is sand and water

**Suspension** - looks like a **solution** when **stirred** but **settles** a bit of water (suspensions). Also it is a mixture - gaseous mixture where the small solid particles of a substance are spread throughout a liquid without settling.

**Unsaturated** - solutions with solubility chart when the temperature would be enough to make a substance dissolve but at a greater amount of solute solubility **less solute dissolves than can** (can't all go into there)

**Saturated** - solutions with solubility chart when the temperature-solubility relationship is at the line in other words a solution with solute has **reached** at the maximum possible amount.

**Supersaturated** - contains **more than the max. amount** of solute that can be dissolved at the temperature on the solubility chart this would mean the grams of solute would go above the line.

## 2. SOLUBILITY VS. TEMPERATURE CHART



Above the line - **supersaturated**

At the line - **saturated**

Under the line - **unsaturated**

## 3. METHODS OF SEPARATION

**Filtration** - separate particles by **size**

**Distillation** - separate by **boiling point** - where two liquids to be separated.

**Evaporation** - where one to remove a **liquid from a solid** by **transforming** the liquid to a **gaseous state**.

**Centrifugation** - separating molecules by their **densities** by spinning them apart.

**Decanting** - process of separating insoluble or suspension solutions by **pouring** away the liquid from the solid.

**Chromatography** - separation by the **rates** at which components of different solutions move through a medium (ex. gel electrophoresis)

# Acid And Bases Study Guide

**KJ Lindholm-Leary**



## Acid And Bases Study Guide:

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