

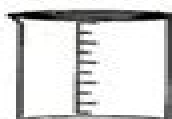
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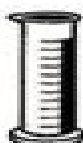
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23. Based on the description of Experiment 1, which of the following diagrams best shows the apparatus that was used to dispense the titrant?

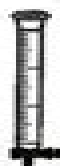
A.



C.



B.



D.



24. Which of the following statements describes a difference between Experiments 1 and 3?

- F. In Experiment 1, the analyte was stirred; in Experiment 3, the analyte was not stirred.
- G. In Experiment 1, the titrant was stirred; in Experiment 3, the titrant was not stirred.
- H. The analyte in Experiment 1 was different from the analyte in Experiment 3.
- J. The titrant in Experiment 1 was different from the titrant in Experiment 3.

25. *Hydrazoic acid* is a weaker acid than is formic acid but is a stronger acid than is acetic acid. Based on Table 1, the pK_a of hydrazoic acid is:

- A. less than 3.2.
- B. between 3.2 and 3.7.
- C. between 3.7 and 4.8.
- D. greater than 4.8.

26. In Experiment 2, the solution in the beaker was neutral when the volume of NaOH added was at a value between:

- F. 9 mL and 11 mL.
- G. 13 mL and 15 mL.
- H. 15 mL and 17 mL.
- J. 19 mL and 21 mL.

27. A student claimed that the acid in UAX is pyruvic acid. Based on Table 1 and Figure 1, this claim is *incorrect* because the pK_a of pyruvic acid is:

- A. less than was the pH at the equivalence point.
- B. greater than was the pH at the equivalence point.
- C. less than was the pH at the halfway point.
- D. greater than was the pH at the halfway point.

28. Suppose Experiment 3 is repeated, but the concentration of the NaOH solution is 0.200 mole/L. Will the volume of NaOH added at the halfway point and at the equivalence point be less than, greater than, or equal to the corresponding results shown in Figure 3?

	halfway point	equivalence point
F.	less	less
G.	greater	less
H.	greater	greater
J.	equal	equal

Act Science Practice Test

Young Scholars Circle

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