

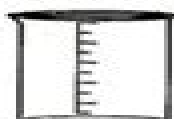
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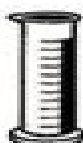
4

23. Based on the description of Experiment 1, which of the following diagrams best shows the apparatus that was used to dispense the titrant?

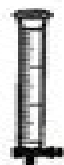
A.



C.



B.



D.



24. Which of the following statements describes a difference between Experiments 1 and 3?

- F. In Experiment 1, the analyte was stirred; in Experiment 3, the analyte was not stirred.
- G. In Experiment 1, the titrant was stirred; in Experiment 3, the titrant was not stirred.
- H. The analyte in Experiment 1 was different from the analyte in Experiment 3.
- J. The titrant in Experiment 1 was different from the titrant in Experiment 3.

25. Hydrazoic acid is a weaker acid than is formic acid but is a stronger acid than is acetic acid. Based on Table 1, the pK_a of hydrazoic acid is:

- A. less than 3.2.
- B. between 3.2 and 3.7.
- C. between 3.7 and 4.8.
- D. greater than 4.8.

26. In Experiment 2, the solution in the beaker was neutral when the volume of NaOH added was at a value between:

- F. 9 mL and 11 mL.
- G. 13 mL and 15 mL.
- H. 15 mL and 17 mL.
- J. 19 mL and 21 mL.

27. A student claimed that the acid in UAX is pyruvic acid. Based on Table 1 and Figure 1, this claim is incorrect because the pK_a of pyruvic acid is:

- A. less than was the pH at the equivalence point.
- B. greater than was the pH at the equivalence point.
- C. less than was the pH at the halfway point.
- D. greater than was the pH at the halfway point.

28. Suppose Experiment 3 is repeated, but the concentration of the NaOH solution is 0.200 mole/L. Will the volume of NaOH added at the halfway point and at the equivalence point be less than, greater than, or equal to the corresponding results shown in Figure 3?

	halfway point	equivalence point
F.	less	less
G.	greater	less
H.	greater	greater
J.	equal	equal

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Table of Contents Act Explore Practice Science Tests 8th Grade

1. Understanding the eBook Act Explore Practice Science Tests 8th Grade
 - The Rise of Digital Reading Act Explore Practice Science Tests 8th Grade
 - Advantages of eBooks Over Traditional Books
2. Identifying Act Explore Practice Science Tests 8th Grade
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Act Explore Practice Science Tests 8th Grade
 - User-Friendly Interface
4. Exploring eBook Recommendations from Act Explore Practice Science Tests 8th Grade
 - Personalized Recommendations
 - Act Explore Practice Science Tests 8th Grade User Reviews and Ratings
 - Act Explore Practice Science Tests 8th Grade and Bestseller Lists

5. Accessing Act Explore Practice Science Tests 8th Grade Free and Paid eBooks
 - Act Explore Practice Science Tests 8th Grade Public Domain eBooks
 - Act Explore Practice Science Tests 8th Grade eBook Subscription Services
 - Act Explore Practice Science Tests 8th Grade Budget-Friendly Options
6. Navigating Act Explore Practice Science Tests 8th Grade eBook Formats
 - ePub, PDF, MOBI, and More
 - Act Explore Practice Science Tests 8th Grade Compatibility with Devices
 - Act Explore Practice Science Tests 8th Grade Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Act Explore Practice Science Tests 8th Grade
 - Highlighting and Note-Taking Act Explore Practice Science Tests 8th Grade
 - Interactive Elements Act Explore Practice Science Tests 8th Grade
8. Staying Engaged with Act Explore Practice Science Tests 8th Grade
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Act Explore Practice Science Tests 8th Grade
9. Balancing eBooks and Physical Books Act Explore Practice Science Tests 8th Grade
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Act Explore Practice Science Tests 8th Grade
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Act Explore Practice Science Tests 8th Grade
 - Setting Reading Goals Act Explore Practice Science Tests 8th Grade
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Act Explore Practice Science Tests 8th Grade
 - Fact-Checking eBook Content of Act Explore Practice Science Tests 8th Grade
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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