

Acid Base Equilibria

22 pages

Titrations

- It needs acid and the concentration of the acid is a base
- If the acid/base of **known concentration** is the **titrant**
- If the acid/base of **unknown concentration** the **titrant** is the **analyte** to be added
- If adding base to solution into the conical flask until the solution is **neutralised** - stops to turn into the next point
- Indicator changes are done from the **colour** of the substance as it is required to neutralise the acid with the strong base - colour change is observed from red
- Repeat until you get 1 **concentration** correctly - within 0.1 pH unit

Titration curve

Shows pH against volume added that is reacting

Strong acid - strong base



- If strong acid is in the conical flask and base is being added to it
- If strong base, the graph shows the solution is the **weak acid** flask base - base from the titrant is being added to it
- If weak at pH 7 or there is an **indicator** of **weak acid** in the conical flask
- If flask neutral (7) so no more base **concentration** change

Weak acid - strong base



Bronsted Lowry

Acids - give up electrons
Bases - gain electrons

- If strong acids are mixed with water, hydrogen ions are released. These H⁺ ions then combine with OH⁻ which make the solution acidic



- If HCl is the acid and H₂O is the base so the forward reaction is to react it is the base
- If HCl and H₂ are conjugate acid base pairs
- If HCl and H₂O are conjugate acid base pairs

- If weak acids are in base where an acid is added to it
- If weak bases are mixed with water, hydroxide ions are released



- If weak acidities in a base then there is a larger concentration of hydroxide ions than the acid
- If strong acids and bases react with water, the reaction is irreversible
- If acid base reactions involve transfer of protons

Conjugate pairs

Select the transfer of electrons



- If HCl is the acid and H₂O is the base reaction so it is donating electrons with it to form the acid
- If HCl and H₂O are conjugate acid base pairs
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- If HCl and H₂O are conjugate acid base pairs

Acid base reactions

ionic product of water

Water exists in equilibrium with its ions
Water dissociates into hydroxide and hydronium ions. $[OH^-] = [H_3O^+] = 10^{-7}$

This equation is simplified to $[H^+] \times [OH^-] = 10^{-14}$

Where does K_w come from?

- If water dissociates into its ions so we know the concentration of water has a constant value
- If the resulting H_3O^+ and OH^- to give a new constant value is K_w - the ionic product of water
- If K_w is the water in a solution at a given temperature
- If this value changes if the temperature changes
- If this value changes if the temperature changes

$pH = -\log[H^+]$ in display K_w values are a number with nothing to do with it



Example

What is the pH of a 0.1M NaOH solution?

If each mole of NaOH gives 1 mole of OH⁻ ions so the concentration of OH⁻ ions is about 0.1M and then



Most strong bases dissociate to produce one OH⁻ ion so it is very easy to calculate. This means the concentration of the OH⁻ ions is equal to the concentration of the base

Weak bases

- If weak bases dissociate into
- If equilibrium lies with to the left and the reaction is reversible
- If this means hydroxide ions formed
- If the reaction is reversible



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Table of Contents Acid Baase Equilibria Study Guide Answers

1. Understanding the eBook Acid Baase Equilibria Study Guide Answers
 - The Rise of Digital Reading Acid Baase Equilibria Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Acid Baase Equilibria Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acid Baase Equilibria Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acid Baase Equilibria Study Guide Answers
 - Personalized Recommendations
 - Acid Baase Equilibria Study Guide Answers User Reviews and Ratings

- Acid Baase Equilibria Study Guide Answers and Bestseller Lists
- 5. Accessing Acid Baase Equilibria Study Guide Answers Free and Paid eBooks
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 - Acid Baase Equilibria Study Guide Answers Budget-Friendly Options
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 - Acid Baase Equilibria Study Guide Answers Compatibility with Devices
 - Acid Baase Equilibria Study Guide Answers Enhanced eBook Features
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 - Highlighting and Note-Taking Acid Baase Equilibria Study Guide Answers
 - Interactive Elements Acid Baase Equilibria Study Guide Answers
- 8. Staying Engaged with Acid Baase Equilibria Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acid Baase Equilibria Study Guide Answers
- 9. Balancing eBooks and Physical Books Acid Baase Equilibria Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acid Baase Equilibria Study Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Acid Baase Equilibria Study Guide Answers
 - Setting Reading Goals Acid Baase Equilibria Study Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Acid Baase Equilibria Study Guide Answers
 - Fact-Checking eBook Content of Acid Baase Equilibria Study Guide Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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