

Chemistry: Acids and Bases Study Guide (Ch. 15)

1. Give the definitions of an acid and a base according to the Arrhenius' Definition.
2. Give the definition of an acid and a base according to the bronsted-lowry definition.
3. Which of the definitions is used more often? Why is this?
4. What two substances are created when an acid is mixed with a base?
5. What are the tastes of acids and bases?
6. How does a base feel? How does that compare to the texture of acids?
7. What type of gas is released when an acid mixes with a metal (assuming that there is a reaction)?
8. Acids and bases break down into charge ions called _____ that can conduct electricity. Does strong or weak acid/base conduct more electricity?
9. What is an indicator? What are the 3 types of indicators in the form of paper? What are some indicators in the form of solution?
10. What indicator solution is clear in an acid, but turns pink in a base (pH >8)?
11. What is the ion product (number) for water at 25 °C?
12. What is an amphoteric substance?
13. What determines the strength of an acid? List strong acids and weak acids.
14. How is the strength of a base determined? List strong bases and weak bases.
15. Give the formulas of the following acids: hydrochloric acid, nitrous acid, acetic acid, carbonic acid, and sulfuric acid.
16. What are the names of the following acids? HNO_3 , HF , H_3PO_4 , H_2SO_4 .
17. What color does blue litmus paper turn in an acidic solution? A basic solution?
18. What is the range of pH for an acid or a base of the known concentration?
19. What is a solution with a pH=7 called? Give an example of a solution with this pH.
20. Describe the concept of conjugate acid-base pair. What is the different between acid and conjugate acid? Base and conjugate base? Show an example.
21. What is a buffer solution?
22. What type of solution contains an acid or base of a known concentration in titration?
23. What is the pH of a solution where $[\text{OH}^-]$ is higher than $[\text{H}_3\text{O}^+]$? Give a range.
24. If the $[\text{H}_3\text{O}^+]$ is greater than $[\text{OH}^-]$, is the solution acidic or basic?
25. What is the end point of titration? How is this different from equivalence point?
26. What should the end point of titration be close to?
27. What is a buffer capacity?
28. IF $[\text{H}_3\text{O}^+]$ or $[\text{OH}^-]$ are added beyond the buffer capacity, what will happen to the pH?

Problems

29. If pH is 4, what is pOH?
30. In a solution of water, the concentration of OH^- ions is $1.0 \times 10^{-6}\text{M}$. What is the concentration of the hydronium ions?
31. The concentration of OH^- ions in a solution is $3.4 \times 10^{-3}\text{M}$. Is the solution acidic basic or neutral? What is the pH of the solution?
32. A volume of 37 mL of 0.36 M KCN neutralizes a 75-mL sample of HClO solution (acid). What is the concentration of HClO ?

Acid And Bases Study Guide Ch 15

Michael Rajnik



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