

Table 18.2

Solubility Product Constants (K_{sp}) at 25°C

Salt	K_{sp}	Salt	K_{sp}	Salt	K_{sp}
Halides		Sulfates		Hydroxides	
AgCl	1.8×10^{-10}	PbSO ₄	6.3×10^{-7}	Al(OH) ₃	3.0×10^{-34}
AgBr	5.0×10^{-13}	BaSO ₄	1.1×10^{-10}	Zn(OH) ₂	3.0×10^{-16}
AgI	8.3×10^{-17}	CaSO ₄	2.4×10^{-5}	Ca(OH) ₂	6.5×10^{-6}
PbCl ₂	1.7×10^{-5}	Sulfides		Mg(OH) ₂	7.1×10^{-12}
PbBr ₂	2.1×10^{-6}			Fe(OH) ₂	7.9×10^{-16}
PbI ₂	7.9×10^{-9}			Carbonates	
PbF ₂	3.6×10^{-8}				
CaF ₂	3.9×10^{-11}				
Chromates					
PbCrO ₄	1.8×10^{-14}	NiS	4.0×10^{-20}	CaCO ₃	4.5×10^{-9}
Ag ₂ CrO ₄	1.2×10^{-12}	CuS	8.0×10^{-37}	SrCO ₃	9.3×10^{-10}
		Ag ₂ S	8.0×10^{-51}	ZnCO ₃	1.0×10^{-10}
		ZnS	3.0×10^{-23}	Ag ₂ CO ₃	8.1×10^{-12}
		FeS	8.0×10^{-19}	BaCO ₃	5.0×10^{-9}
		CdS	1.0×10^{-27}		
		PbS	3.0×10^{-28}		

A Solubility Product Constant Lab 17a Answers

**Michael Dzwiniel, Frank Jenkins, Dick
Tompkins, Hans Van Kessel**



A Solubility Product Constant Lab 17a Answers:

The Software Encyclopedia ,1988 **Nelson Chemistry** Michael Dzwiniel, Frank Jenkins, Dick Tompkins, Hans Van Kessel, 1992 Basic resource for Senior High Chemistry 20 30 ca 1993 2007

A Solubility Product Constant Lab 17a Answers Book Review: Unveiling the Power of Words

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