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web a reaction and its experimental rate law are represented below $2 \text{ICl} + \text{H}_2 + \text{O}_2 \rightarrow 2 \text{HCl} + \text{H}_2\text{O}$ rate $k [\text{ICl}]^2 [\text{H}_2]$ a chemist proposes a mechanism for the reaction that is consistent with the rate law the mechanism has two elementary steps and the first step is slow compared to the second

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