#### This activity will help you become skilled at balancing chemical equations.

- 1. Cut out the squares that the bottom of the paper.
- Decide whether or not each equation is balanced. If it is not, decide how to balance it, pasting the small squares as needed into the spaces.
- 3. Some spaces will not require a number.
- S<sub>8</sub> + \_\_\_ O<sub>2</sub> → \_\_\_ SO<sub>3</sub>
- 2. \_\_\_\_ Na + \_\_\_\_ H<sub>2</sub>O → \_\_\_\_ NaOH + \_\_\_\_ H<sub>2</sub>
- 3. \_\_\_C<sub>10</sub>H<sub>16</sub> + \_\_\_Cl<sub>2</sub> → \_\_\_C + \_\_\_HCl
- 4. \_\_\_ Si<sub>2</sub>H<sub>3</sub> + \_\_\_ O<sub>2</sub> → \_\_\_ SiO<sub>2</sub> + \_\_\_ H<sub>2</sub>O
- FeS<sub>2</sub> + \_\_\_ O<sub>2</sub> → \_\_\_ Fe<sub>2</sub>O<sub>3</sub> + \_\_\_ SO<sub>2</sub>
- 6. \_\_\_ Fe<sub>2</sub>O<sub>2</sub> + \_\_\_ H<sub>2</sub>  $\rightarrow$  \_\_\_ Fe + \_\_\_ H<sub>2</sub>O
- 7. \_\_\_ H<sub>2</sub>O<sub>2</sub> → \_\_\_ H<sub>2</sub>O + \_\_\_ O<sub>2</sub>
- 8. \_\_\_ SiO<sub>2</sub> + \_\_\_ HF → \_\_\_ SiF<sub>4</sub> + \_\_\_ H<sub>2</sub>O
- 9. \_\_\_ K + \_\_\_ Br<sub>2</sub> → \_\_\_ KBr
- 10. \_\_\_S<sub>8</sub> + \_\_\_ F<sub>2</sub> → \_\_\_ SF<sub>6</sub>
- 4 11 2 2 2 8 2 16 2 6
  - 8 8 4 11 2 2 2 2 2 8 10
- 12 2 2 4 2 24 8

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**M Planty** 

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