# AP CHEMISTRY MULTIPLE CHOICE OUESTIONS Unit 8 - Introduction to Equilibrium

 $N_2O_4(g) \rightleftharpoons 2NO_2(g) + \frac{1}{2}O_2(g)$ 1.0

> The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the ourilibrium constant for the following reaction BE 25 YEAR

- CAU (95)2
- GE1 95

3...  $C_*H_*(g) + H_*(g) \xrightarrow{\text{plainers}} C_*H_*(g)$ 

C<sub>2</sub>H<sub>2</sub>(g) is reduced by H<sub>2</sub>(g) in the presence of a solid platinum cutalyst, as represented by the equation above. Factors that could affect the rate of the reaction include which of the following?

- Changes in the partial pressure of H<sub>2</sub>(g).
- II. Changes in the particle size of the platinum.
- III. Changes in the temperature of the reaction. SALES SECTION
- (A) III only
- (B) I and II only
- (C) I and III only
- (D) III and III only (Ellip L. III., amed IIII.

shut+ 2 Bao(4) = 2 Bao(4) + O(4) AH = 162 kumol en dotramic

A sealed rigid vessel contains BaO<sub>2</sub>(x) in equilibrium with BaO(s) and O<sub>2</sub>(g) as represented by the equation above. Which of the following changes will increase the amount of BaO<sub>7</sub>(s) in the vessel?

- (A) Removing a small amount of O<sub>2</sub>(x) | K
- (B) Removing a small amount of BaO(x) X
- (C) Adding He gas to the vessel X
- (D) Lowering the temperature \

- For which of the following reaction mixtures at equilibrium would reducing the volume of the container at constant temperature cause the equilibrium to shift toward the products? [ NOVAL (A) SrCO<sub>2</sub>(t) = SrO(1) + CO<sub>2</sub>(g)
  - (B) C(c) + CO(c) = 2 CO(c) (C) CO(g) + H<sub>2</sub>O(g) = CO<sub>2</sub>(g) + H<sub>2</sub>(g)
  - DEROUS = POLED + CILED DEROUS = NOVED towards fewer Shift towards gas

X(g) + Y(g) = 2Z(g)

When 4.00 mot each of X(g) and Y(g) are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of Z(g) is produced. What is the value of the equilibrium constant,  $K_c$  ?

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The synthesis of  $CH_2OH(g)$  from CO(g) and  $H_2(g)$ is represented by the equation above. The value of K for the reaction at 483 K is 14.5.

increase heat shifts le Which of the following explains the effect on the equilibrium constant. K,., when the temperature of the reaction system is increased to 650 K?

- (A) K, will increase because the activation. energy of the forward reaction increases. more than that of the reverse reaction.
- (B) K, will increase because there are more
- reactant molecules than product molecules. (C) K, will decrease because the reaction is exothermic.
- (D) K, is constant and will not change.

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Richard H. Langley, John Moore

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