

Name _____ Date _____ Class _____

CHAPTER 15 **STUDY GUIDE FOR CONTENT MASTERY**

Solutions

Section 15.1 What are solutions?

In your textbook, read about the characteristics of solutions.

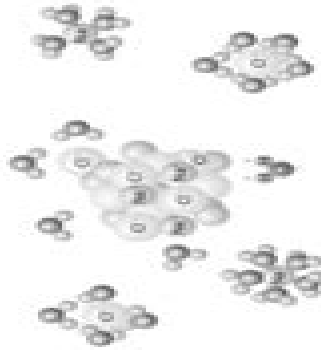
Use each of the terms below just once to complete the passage.

immiscible	liquid	soluble	solution
insoluble	miscible	solvent	solvent

Air is air! (1) solution of oxygen gas dissolved in nitrogen gas. The oxygen is air is the (2) solvent and nitrogen is the (3) solvent. Because oxygen gas dissolves in a solvent, oxygen gas is air! (4) soluble substances. A substance that does not dissolve is (5) insoluble. (6) liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are (7) miscible. However, if one liquid is insoluble in another, the liquids are (8) immiscible.

Read about solvation in aqueous solutions in your textbook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.



8. Hydration is solvation in which the solvent is water. What is solvation?

Solvation is the process of surrounding solute particles with solvent particles to form a solution.

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Chemistry: Matter and Change • Chapter 15

85

Name _____ Date _____ Class _____

CHAPTER 15 **STUDY GUIDE FOR CONTENT MASTERY**

Section 15.1 continued

10. As sodium chloride dissolves in water, what happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

11. Explain the attraction of the water molecules around the sodium ions and chloride ions.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

12. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.

13. List three ways that the rate of solvation may be increased.

stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent

In your textbook, read about heat of solution, solubility, and factors that affect solubility.

For each statement below, write true or false.

true 14. The overall energy change that occurs when a solution forms is called the heat of solution.

false 15. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.

false 16. Solvation continues as long as the solvation rate is less than the crystallization rate.

true 17. In a saturated solution, solvation and crystallization are in equilibrium.

true 18. Additional solute can be dissolved in an unsaturated solution.

true 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

86 Chemistry: Matter and Change • Chapter 15

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Answer Key Chapter 22 Chemistry Study Guide For Content Mastery

Arthur Beiser



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Chemistry Steven S. Zumdahl, Susan A. Zumdahl, 2000 Zumdahl's conceptual approach to problem solving early and thorough treatment of reactions and integration of descriptive chemistry and chemical principles throughout are key elements to the success of this bestseller By encouraging students to understand qualitative concepts before considering quantitative problems Chemistry helps students build a solid understanding of topics and avoid common misconceptions The text also includes new Chemical Impact boxes that present everyday applications of chemistry plus over 200 redesigned illustrations with greater clarity and emphasis on molecular level illustrations **Resources in Education**, 1984 Serves as an index to Eric reports microform *Merrill Chemistry* Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

Essentials of Ecology George Tyler Miller, 2006-09 G Tyler Miller's worldwide bestsellers have evolved right along with the changing needs of your diverse student population Focused specifically on energizing and engaging all your students Miller and new contributor Scott Spoolman have been at work scrutinizing every line enhancing clarifying and streamlining to reduce word density as well as updating with the very latest environmental news and research The resulting texts are shorter clearer and so engaging that your students will actually want to read their assignments The ideal alternative to ecology texts that tend to be too difficult for non majors this succinct 13 chapter full color textbook covers scientific principles and concepts ecosystems evolution biodiversity population ecology and more New to this edition for instructors is PowerLecture a one stop shop for lecture prep that includes everything you need to create dynamic lectures all in one place

The Mainstream of Physics Arthur Beiser, 1962 *Resources in Education*, 1984 **Popular Mechanics**, 2000-01 Popular Mechanics inspires instructs and influences readers to help them master the modern world Whether it's practical DIY home improvement tips gadgets and digital technology information on the newest cars or the latest breakthroughs in science PM is the ultimate guide to our high tech lifestyle *Glencoe Chemistry* Glencoe/McGraw-Hill, 2002 Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry Denise Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts After a brief review of each section's critical ideas students are taken through stepped out worked examples try it yourself examples and chapter quizzes all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text **Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick

access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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Table of Contents Answer Key Chapter 22 Chemistry Study Guide For Content Mastery

1. Understanding the eBook Answer Key Chapter 22 Chemistry Study Guide For Content Mastery
 - The Rise of Digital Reading Answer Key Chapter 22 Chemistry Study Guide For Content Mastery
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Chapter 22 Chemistry Study Guide For Content Mastery
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Chapter 22 Chemistry Study Guide For Content Mastery
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Chapter 22 Chemistry Study Guide For Content Mastery
 - Personalized Recommendations
 - Answer Key Chapter 22 Chemistry Study Guide For Content Mastery User Reviews and Ratings
 - Answer Key Chapter 22 Chemistry Study Guide For Content Mastery and Bestseller Lists
5. Accessing Answer Key Chapter 22 Chemistry Study Guide For Content Mastery Free and Paid eBooks
 - Answer Key Chapter 22 Chemistry Study Guide For Content Mastery Public Domain eBooks
 - Answer Key Chapter 22 Chemistry Study Guide For Content Mastery eBook Subscription Services
 - Answer Key Chapter 22 Chemistry Study Guide For Content Mastery Budget-Friendly Options
6. Navigating Answer Key Chapter 22 Chemistry Study Guide For Content Mastery eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key Chapter 22 Chemistry Study Guide For Content Mastery Compatibility with Devices
 - Answer Key Chapter 22 Chemistry Study Guide For Content Mastery Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key Chapter 22 Chemistry Study Guide For Content Mastery
 - Highlighting and Note-Taking Answer Key Chapter 22 Chemistry Study Guide For Content Mastery
 - Interactive Elements Answer Key Chapter 22 Chemistry Study Guide For Content Mastery
8. Staying Engaged with Answer Key Chapter 22 Chemistry Study Guide For Content Mastery

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key Chapter 22 Chemistry Study Guide For Content Mastery
9. Balancing eBooks and Physical Books Answer Key Chapter 22 Chemistry Study Guide For Content Mastery
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Chapter 22 Chemistry Study Guide For Content Mastery
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Key Chapter 22 Chemistry Study Guide For Content Mastery
- Setting Reading Goals Answer Key Chapter 22 Chemistry Study Guide For Content Mastery
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key Chapter 22 Chemistry Study Guide For Content Mastery
- Fact-Checking eBook Content of Answer Key Chapter 22 Chemistry Study Guide For Content Mastery
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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