

TESTS

EXPLORE CHEMISTRY WITH CHEMISTRY

DESCRIBING CHEMICAL REACTIONS

SOLUTIONS TO THE MODULE 7 REVIEW QUESTIONS

1. Elements cannot undergo decomposition reactions because elements have nothing to decompose into.
2. a. This equation represents a synthesis reaction because 2 elements come together to produce 1 compound.
b. This equation represents a formation reaction because 2 elements come together to produce 1 compound.
3. None of these
4. None of these
5. This equation represents a decomposition reaction because a single compound is breaking down into its elements.
6. The only equation that does not represent a formation reaction is (b). The equation for (b) is a formation equation because it starts with 2 reactants and makes only 1 in the end. Remember, reactions do not have to start with elements in order to be formation equations; we just assume they do when we write them ourselves.
7. The only equation that does not represent a combustion reaction is (b). The equation for (b) is a decomposition equation. Even though it produces CO_2 and H_2O , it does not have O_2 as a reactant, so it is not combustion.
8. Complete combustion produces carbon dioxide and water. Incomplete combustion produces carbon monoxide and water or carbon and water.

...ment units, and 4 g for the answer. If the units, this is the same as g/mol, but we will go on to see that these are conflicting units. The units are given in kg. That, before we solve the units kg to g.

... can use the equation for density to solve for volume.

... each answer, minus 2 g for every wrong element found and minus 1 g for every wrong element found. The maximum score for this test is 100 g.

The equation is balanced.

... $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$...

... 2 g for the correct equation, 4 g for the equation, and 4 g for the units following it. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$...

... 4 g for the correct element compounds, 4 g for the equation, and 4 g for the units following it. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$...

... $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$...

... 4 g for the correct equation is balanced, then the same number of atoms appear on both sides. They would appear the same in the same, whether we are adding, subtracting, or multiplying. That's another way of saying the law of conservation.

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Jiyuan Zhang



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