

Задача 11
Слабые и сред. кислоты

1) Дано:

$$C_{\text{HNO}_2} = 0,01 \text{ м}$$

$$K_a = 7,9 \cdot 10^{-4}$$

$$pH = ?$$

$$pH = -\lg [H^+] = -\lg 2,8 \cdot 10^{-5} = 4,55$$

Решение:

Для слабой кислоты:

$$[H^+] = \sqrt{K_a \cdot C_{\text{HNO}_2}}$$

$$= \sqrt{7,9 \cdot 10^{-4} \cdot 10^{-2}} = 2,8 \cdot 10^{-5} \text{ моль/л}$$

$$= 4,55$$

2) Дано:

$$K_{\text{PrOH}_2} = 6,3 \cdot 10^{-5}$$

$$pH = ?$$

ξ - гидролизная степень

$$\xi = \sqrt{\frac{K_{\text{Pr}}}{C}} = \sqrt{\frac{6,3 \cdot 10^{-5}}{4}} = 1,25 \cdot 10^{-2} \text{ моль/л}$$

$$[OH^-] = \frac{\xi}{2} = \frac{1,25 \cdot 10^{-2}}{2} = 6,25 \cdot 10^{-3} \text{ моль/л}$$

$$[H^+] = \frac{10^{-14}}{[OH^-]} = \frac{10^{-14}}{6,25 \cdot 10^{-3}} = 1,6 \cdot 10^{-12}$$

$$pH = -\lg [H^+] = -\lg 1,6 \cdot 10^{-12} = 12 - 0,2 = 11,8$$

Решение:

$$C_{\text{PrOH}_2} = C_{\text{H}_2\text{PO}_4^{2-}} = 2,5 \cdot 10^{-2}$$

$$K_{\text{Pr}} = [Ca^{2+}][OH^-]^2$$

$$= 6 \cdot (2,5)^2 = 37,5$$

3) Дано:

$$V_{\text{HNO}_2} = V_{\text{HNO}_3}$$

$$C_{\text{HNO}_2} = 0,1 \text{ м}$$

$$V_{\text{PrOH}_2} = 1 \text{ см}^3$$

$$C_{\text{PrOH}_2} = 0,01 \text{ м}$$

$$K_{\text{PrOH}_2} = 4,6 \cdot 10^{-5}$$

Решение:

полн. самодиссоции:

$$C_{\text{HNO}_2} = 0,1 \frac{\text{г}}{\text{л}} = 0,05 \text{ моль/л}$$

$$C_{\text{PrOH}_2} = 0,01 \frac{\text{г}}{\text{л}} = 0,005 \text{ моль/л}$$

Для гидролизующей кислоты:

$$C_{\text{H}^+} = C_{\text{HNO}_2} = 0,05 \text{ моль/л}$$

$$C_{\text{PrOH}_2} = C_{\text{H}_2\text{PO}_4^{2-}} = 0,005 \text{ моль/л}$$

$$[Pr^+][Cl^-]^2 = 4,6 \cdot 10^{-5} = 10^{-4,34}$$

$$PrCl_2 \rightleftharpoons Pr^{2+} + 2Cl^-$$

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Ashy Kumren, Chwee Neo Lee



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