

AP CHEMISTRY MULTIPLE CHOICE QUESTIONS Unit 8 – Introduction to Equilibrium



The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the equilibrium constant for the following reaction at 25°C?



Reversed
+ Doubled

(A) 95²

(B) 95

(C) (95)^{1/2}

(D) $\frac{1}{95}$

(E) $\frac{1}{(95)^2}$



$C_2H_4(g)$ is reduced by $H_2(g)$ in the presence of a solid platinum catalyst, as represented by the equation above. Factors that could affect the rate of the reaction include which of the following?

- I. Changes in the partial pressure of $H_2(g)$
- II. Changes in the particle size of the platinum catalyst
- III. Changes in the temperature of the reaction system

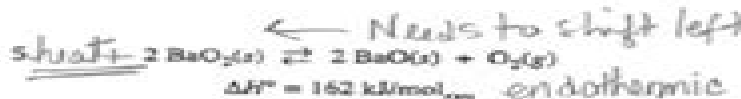
(A) III only

(B) I and II only

(C) I and III only

(D) II and III only

(E) I, II, and III



A sealed rigid vessel contains $BaO_2(s)$ in equilibrium with $BaO(s)$ and $O_2(g)$ as represented by the equation above. Which of the following changes will increase the amount of $BaO_2(s)$ in the vessel?

(A) Removing a small amount of $O_2(g)$ $R \times$

(B) Removing a small amount of $BaO(s)$ X

(C) Adding He gas to the vessel X

(D) Lowering the temperature L

2. For which of the following reaction mixtures at equilibrium would reducing the volume of the container at constant temperature cause the equilibrium to shift toward the products?



increase P°
shift towards fewer
+ make gas



When 4.00 mol each of $X(g)$ and $Y(g)$ are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of $Z(g)$ is produced. What is the value of the equilibrium constant, K_c ?

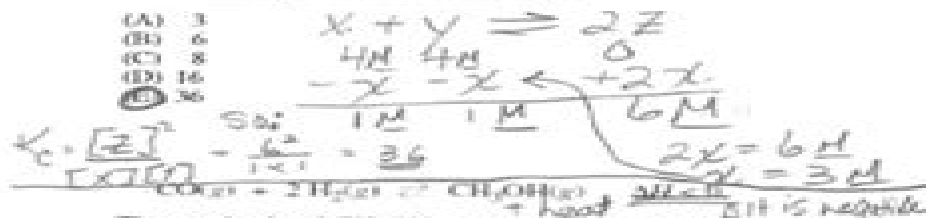
(A) 3

(B) 6

(C) 8

(D) 16

(E) 36



6. The synthesis of $CH_3OH(l)$ from $CO_2(g)$ and $H_2(g)$ is represented by the equation above. The value of K_c for the reaction at 483 K is 14.5.

increase heat shifts left, decrease K_c
Which of the following explains the effect on the equilibrium constant, K_c , when the temperature of the reaction system is increased to 650 K?

(A) K_c will increase because the activation energy of the forward reaction increases more than that of the reverse reaction.

(B) K_c will increase because there are more reactant molecules than product molecules.

(C) K_c will decrease because the reaction is exothermic.

(D) K_c is constant and will not change.

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web step 1 inspect parts critical inspection the output shaft and lower planetary assemblies and related hardware are shown here remove the low planetary carrier from the output shaft carrier assembly to inspect the pinions torrington bearings and replace the bushing step 2 remove snap ring

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web step 1 remove modulator remove the bolt for the transmission modulator retainer remove the modulator and the modulator valve a strong magnet usually pulls the valve from the case if it doesn't want to come out you can use a small flat tip screwdriver to help it out of the case after the valve body is removed step 2 remove governor