

Basic Chemistry Study Guide 2

1. Hydrogen bonds

1. Form between two polar molecules **covalent bonds**
2. A hydrogen's δ^+ region is attracted to a δ^- region in another molecule.
3. Form between water molecules (Where are the δ^+ and δ^- regions in a water molecule? **hydrogen bonds, at the top of the symbol**)
4. Each hydrogen bond is very weak (can be easily formed and broken).
5. There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life
6. One of these properties is the ability to be a solvent for other polar molecules.
7. Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

2. Solutions, solvents, and solutes

1. Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
2. Water is located inside and outside of our cells.
3. Body fluids are solutions
4. Define the terms **solution, solvent, and solute**. **Solvent: a substance capable of dissolving other polar molecules and ionic bonds, solutes: molecules dissolved in the solvent, solution: homogeneous mixture**

3. How substances dissolve in water

1. When a solute dissolves, it becomes surrounded by water molecules.
2. For a solute to dissolve, it must attract water molecules.
3. What kinds of substances attract water molecules? **Sugars, nucleic acids, and amino acids**
4. When a solute dissolves it forms hydrogen bonds with water molecules.
5. Define the terms **ion, cation, and anion**. **ion: atom with a charge, cation: positively charged ion, anions: negatively charged ion**
6. Define the terms **hydrophilic** and **hydrophobic** **Hydrophilic: attracted to water, hydrophobic: afraid of water**
7. What types of molecules are hydrophilic and what types are hydrophobic? **Nonpolar hydrophobic, polar hydrophilic**

Basic Chemistry Study Guide

Karen C. Timberlake, Mark Quirie



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1. Understanding the eBook Basic Chemistry Study Guide
 - The Rise of Digital Reading Basic Chemistry Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Basic Chemistry Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Basic Chemistry Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Basic Chemistry Study Guide
 - Personalized Recommendations
 - Basic Chemistry Study Guide User Reviews and Ratings
 - Basic Chemistry Study Guide and Bestseller Lists
5. Accessing Basic Chemistry Study Guide Free and Paid eBooks
 - Basic Chemistry Study Guide Public Domain eBooks
 - Basic Chemistry Study Guide eBook Subscription Services
 - Basic Chemistry Study Guide Budget-Friendly Options
6. Navigating Basic Chemistry Study Guide eBook Formats

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- Basic Chemistry Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Basic Chemistry Study Guide
 - Highlighting and Note-Taking Basic Chemistry Study Guide
 - Interactive Elements Basic Chemistry Study Guide
- 8. Staying Engaged with Basic Chemistry Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Basic Chemistry Study Guide
- 9. Balancing eBooks and Physical Books Basic Chemistry Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Basic Chemistry Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Basic Chemistry Study Guide
 - Setting Reading Goals Basic Chemistry Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Basic Chemistry Study Guide
 - Fact-Checking eBook Content of Basic Chemistry Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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