

# LAB MANUAL CHEMISTRY CLASS XII

Easy!

## Experiment 1

Aim: To prepare 250 ml of 0.1N of Mohr's salt

Theory

Molecular mass of Mohr's salt

$\text{FeSO}_4 \cdot (\text{NH}_4)_2\text{SO}_4 \cdot 6\text{H}_2\text{O}$  is 392 g

Hence, for preparing 250 ml of 0.1N Mohr's salt

solution 3.92 g of salt is required

for preparing 250 ml of 0.1N Mohr's salt  
solution  $\frac{3.92 \times 250 \times 100}{1000} = 9.9 \text{ g}$  is required

Apparatus required: watch glass, weights, 250 ml beaker, glass rod, wash bottle.

Chemicals required: Mohr's salt, conc.  $\text{H}_2\text{SO}_4$ , distilled water.

Procedure

1. Record weight of clean dry watch glass

2. Weigh and record exactly 9.9 g of Mohr's salt crystals

# Answer To Hands On Lab Manual Chemistry

**JR Anderson**



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