

Experiment 25

Advance Study Assignment: pH, Its Measurement and Applications

1. A solution of a weak acid was tested with the indicators used in this experiment. The colors observed were as follows:

Methyl violet	violet > 1.6	Congo red	violet < 3.0
Thymol blue	orange > 2.8	Bromocresol green	yellow < 4.0
Methyl yellow	red < 4.0		

What is the approximate pH of the solution?

$2.5 < \text{pH} < 3.0$

approx 3!

2. An aqueous solution of NH_4^+ has a pH of 11.6. The ammonia (NH_3) molecule is the conjugate base of the NH_4^+ (called "ammonium") ion. Write the net ionic equation for the reaction that makes an aqueous solution of NH_4^+ basic. (Eq. 5.)



3. The pH of a 0.10 M HOBr solution is 4.8.

- a. What is $[\text{H}^+]$ in that solution?

$$1.58 \times 10^{-5} \text{ M}$$

- b. What is $[\text{OBr}^-]$? What is $[\text{HOBr}]$? (Where do the H^+ and OBr^- ions come from?)

$$0.10 \text{ M} - 1.58 \times 10^{-5} \text{ M} \approx 0.10 \text{ M}$$

$$\text{_____ M}; \text{_____ M}$$

- c. What is the value of K_a for HOBr? What is the value of $\text{p}K_a$?

$$\text{_____}$$

Advance Study Assignment Experiment 25 Answers

Scott C. Dulebohn



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